# Chemical Equilibrium:

Equilibrium (K) vs Kinetics (k):	Law of Mass Action:
$A \stackrel{.}{\rightharpoonup} B$	jA+kB=IC+mD
$Rate_{Forward} = Rate_{Reverse}$	The Equilibrium Constant:
$k_1[A] = k_{-1}[B]$ $\frac{k_1}{K_{-1}} = \frac{[B]}{[A]} = K_{eq}$	$K_{c} = \frac{[C]^{l}[D]^{m}}{[A]^{l}[B]^{k}} K_{P} = \frac{ P_{C} ^{l} P_{D} ^{m}}{ P_{A} ^{l} P_{B} ^{k}}$ Note: Solids and liquids are not included in the equilibrium expression. They receive a value of 1.
Reaction Adjustments:	Equilibrium Partial Pressure From Kc:
$A+B \rightarrow C K$	jA+kB=IC+mD
$2A + 2B \rightarrow 2C K^2$	$K_p = K_C  RT ^{\Delta n} R = 0.08206 \frac{L * atm}{mod * K}$
$\frac{1}{2}A + \frac{1}{2}B \rightarrow \frac{1}{2}CK^{1/2}$	$\Delta n =  I+m  - (j+k)$
$C = A + B \frac{1}{K}$	$T_K = T_C + 273.15$
Additive Reactions:	The Reaction Quotient Q:
$A+B \rightarrow C K_1$ $C+D \rightarrow E K_2$	$Q = \frac{[C]'[D]^n}{[A]'[B]^k}$
$A+B+D \rightarrow E k = k_1 \cdot k_2$	Q → Initial Concentrations K → Equilibrium Concentrations
The Position of Equilibrium:	The Direction of the Reaction:
<ul> <li>(a) If K &gt; &gt; 1, Reaction = Product</li> <li>Favored.</li> <li>(b) If 0 &lt; K &lt; &lt; 1, Reaction = Reactant</li> <li>Favored.</li> </ul>	<ol> <li>If Q &lt; K, the reaction will shift to the Right.</li> <li>If Q &gt; K, the reaction will shift to the Left.</li> <li>If Q = K, the system is at equilibrium.</li> </ol>
Le Chatelier's Principle:	Quadratic Equations:
If a change is imposed on a system at equilibrium, the reaction will shift in a direction that tends to reduce that change.	$a x^2 + bx + c = 0$ $x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$

# **Chem Fiesta Equilibrium Equations**

Hossein Ghadiri, Calvin Rose

### **Chem Fiesta Equilibrium Equations:**

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