

Chapter 7 – Chemical Formulas and Chemical Compounds

7-1 Chemical Names and Formulas

I. Significance of a Chemical Formula

A. Molecular formulas

1. Number of atoms of each element in one molecule of a compound
 C_2H_6 = ethane (2 carbon atoms, 6 hydrogen atoms)

B. Ionic Compounds

1. Represents the simplest whole number ratio of the compounds cations and anions



II. Monatomic Ions

A. Monatomic ions

1. ions formed from a single atom

1																17	18
H ⁺	2											13	14	15	16	H ⁻	NOBLE GASES
Li ⁺		3	4	5	6	7	8	9	10	11	12			N ³⁻	O ²⁻	F ⁻	
Na ⁺	Mg ²⁺	Transition metals										Al ³⁺		P ³⁻	S ²⁻	Cl ⁻	
K ⁺	Ca ²⁺				Cr ³⁺	Mn ²⁺	Fe ²⁺ Fe ³⁺	Co ²⁺	Ni ²⁺	Cu ⁺ Cu ²⁺	Zn ²⁺				Se ²⁻	Br ⁻	
Rb ⁺	Sr ²⁺									Ag ⁺	Cd ²⁺		Sn ²⁺		Te ²⁻	I ⁻	
Cs ⁺	Ba ²⁺								Pt ²⁺	Au ⁺ Au ³⁺	Hg ²⁺ Hg ¹⁺		Pb ²⁺	Bi ³⁺			

B. Naming Monatomic Ions

1. Monatomic cations are
 - a. Identified by the element's name
2. Monatomic anions
 - a. Drop the ending of the element name
 - b. Add an "-ide" ending

III. Binary Ionic Compounds

A. Binary Compounds

1. Compounds composed of two different elements

B. Writing Formulas for Binary Ionic Compounds

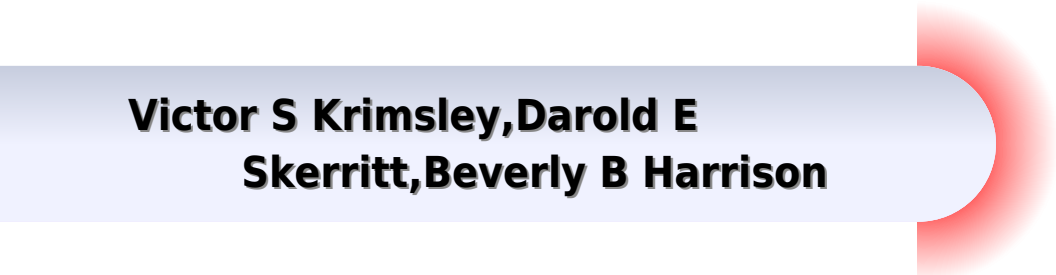
1. Write the symbols for the ions side by side. ALWAYS write the cation first!
2. Cross over the charges by using the absolute value of each ion's charge as the subscript for the other ion
3. Check that the subscripts are in smallest whole number ratio

C. Naming Binary Ionic Compounds from Their Formulas

1. Name the cation.
2. Name the anion.

Chem Worksheet 8

**Victor S Krimsley, Darold E
Skerritt, Beverly B Harrison**



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Boiler Technician 3 & 2 Ronald E. Allen, 1992

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Engineering Administration Jack Garrett (Jr.), 1983 **Addison-Wesley**

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Study of Equipment and Methods for Removing Oil from Harbor Waters Pacific Northwest Laboratory, 1969

Site Assessment System (SAS) for the Michigan Priority Ranking System Under the Michigan Environmental Response Act (Act 307, P.A. 1982)., 1983 *I/M&w/Tsts Intro Chem* Victor S Krimsley, Darold E Skerritt, Beverly B Harrison, 1986

Experimental Chemistry Robert J. Artz, 1982

Chemical Process Safety Daniel A. Crowl, Joseph F. Louvar, 2019-03-01 The 1 Process Safety Guide Now Extensively Updated for Current Industrial Processes Systems and Practices Process safety has seen a dramatic consolidation of concepts in the past few years Chemical Process Safety Fourth Edition provides students and working engineers with the understanding necessary to apply these new concepts to safely design and operate any process Long the definitive guide in the field this edition fully reflects major recent advances in process safety technology and practice Readers will find extensive new and updated coverage of relief sizing hazards identification risk assessment and many other topics Several chapters have been completely rewritten and all are substantially modified This textbook includes 50 new problems and solutions mostly in SI units and 25 new case histories Safety culture Preventive and mitigative safeguards The CCPS 20 elements of Risk Based Process Safety RBPS Toxicology industrial hygiene and source models Hazardous material dispersion Fires explosions and concepts for preventing them Chemical reactivity Reliefs and relief sizing Hazards identification and evaluation Risk analysis and assessment including Layer of Protection Analysis LOPA Safety strategies procedures designs case histories and lessons learned Crowl and Louvar link key academic concepts to modern industrial practice making this guide invaluable for all engineering students and for all working engineers Register your product for convenient access to downloads updates and or corrections as they become available See inside book for details

Chemical Engineering

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