

KEY

Equilibrium – LeChatelier and Calculations

Learning Objectives – By the end of this worksheet, you will be able to:

1. Predict the effect on equilibrium and how it shifts from changes in concentration, temperature, and the presence of a catalyst.
2. Calculate the reaction quotient, Q , and use it to determine the direction the reaction will shift to re-establish equilibrium.
3. Interpret graphical data showing how changing things in LO #1 affect equilibrium, concentrations, and how equilibrium is reestablished.
4. Calculate equilibrium concentrations or the equilibrium constant depending on the information provided (ICE tables).

Part I: Q vs K and LeChatelier

1. Time for practice! Consider the reaction:



The equilibrium concentrations were determined to be 1.00 M H_2 , 0.500 M N_2 , and 0.173 M NH_3 at 400°C. Calculate K for this system.

$$K = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3} \rightarrow K = \frac{(0.173)^2}{(0.500)(1)^3} = \boxed{0.0599}$$

2. Consider the reaction:



$$K_{eq} = 1.2 \times 10^{-12} @ 473K$$

- a. We initially start with 0.37 M H_2 , 0.49 M Cl_2 , and 0.0 M HCl at 473K. Calculate Q for this system.

$$Q = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]} \rightarrow Q = \frac{(0)^2}{(0.37)(0.49)} = \boxed{0}$$

- b. For your system, will the reaction shift left or right to re-establish equilibrium?

$0 < 1.2 \times 10^{-12}$ so $Q < K$; the reaction will shift right to make more products

3. The reaction below is conducted in a 1.00 L flask. The equilibrium concentrations were determined to be 0.500 M H_2 , 0.400 M I_2 , and 1.60 M HI .



- a. Write the law of mass action and calculate K for this system.

$$K = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]} \rightarrow K = \frac{(1.60)^2}{(0.5)(0.4)} = \boxed{12.8}$$

- b. If we add 1.00 mol of H_2 to the system after it has reached equilibrium, will the system still be at equilibrium? Explain your choice. To help your explanation, calculate Q below.

No adding reactant changes the ratio of products to reactants. we aren't at equilibrium anymore.

$$Q = \frac{(1.60)^2}{(1.500)(0.400)} = \boxed{4.27}$$

$Q \neq K$! If Q doesn't equal K, we aren't at equilibrium

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