

Student Name: \_\_\_\_\_ Student ID # \_\_\_\_\_ Section # \_\_\_\_\_

1 atm = 760 mm Hg = 101,325 Pa      0 °C = 273 K      R = 0.08206 (L·atm)/(mol·K) = 8.314 J/(mol·K)

- What is the volume of 21.0 g of nitrogen gas ( $N_2$ ) at STP?  
A) 22.4 L  
B) 11.2 L  
C) 16.8 L  
D) 33.6 L  
E) none of the above
- What is the equivalent pressure of 1520 torr in units of atm?  
A) 203,000  
B) 380  
C) 2.00  
D) 1520  
E) none of the above
- Calculate the density, in g/L, of  $CO_2$  gas at 27°C and 0.50 atm pressure.  
A) 0.89 g/L  
B) 1.12 g/L  
C) 9.93 g/L  
D) 46.0 g/L  
E) 2.17 kg/L
- What is the pressure of a 3.00 L gas vessel that has 18.0 grams of helium at 25 °C?  
A) 147 atm  
B) 36.7 atm  
C) 32.6 atm  
D) 1.81 atm  
E) none of the above
- A mixture of three gases has a total pressure of 1380 mmHg at 298 K. The mixture contains 1.27 mol  $CO_2$ , 3.04 mol  $CO$ , and 1.50 mol Ar. What is the partial pressure of Ar?  
A. 0.258 atm  
B. 301 mmHg  
C. 5345 mmHg  
D. 356 mmHg  
E. 8020 mmHg

# Chem Think Answers Gases

**James Jeans, Sir James Hopwood Jeans**

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*An Introduction to the Kinetic Theory of Gases* James Jeans, Sir James Hopwood Jeans, 1982-10-14 This book can be described as a student's edition of the author's *Dynamical Theory of Gases*. It is written however with the needs of the student of physics and physical chemistry in mind and those parts of which the interest was mainly mathematical have been discarded. This does not mean that the book contains no serious mathematical discussion; the discussion in particular of the distribution law is quite detailed but in the main the mathematics is concerned with the discussion of particular phenomena rather than with the discussion of fundamentals. **Gases**, 1984

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