

Ions and Ionic Compounds (How are ionic compounds formed?)

Model 1: Common charges (by group) on elements when in ionic compounds.

Charge		← variable charge →																			
Grp #																					
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18				
1 H																		2 He			
3 Li	4 Be												5 B	6 C	7 N	8 O	9 F	10 Ne			
11 Na	12 Mg												13 Al	14 Si	15 P	16 S	17 Cl	18 Ar			
19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr				
37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe				
55 Cs	56 Ba	57 La	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn				
87 Fr	88 Ra	89 Ac	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn		114 Fl		116 Lv						

A cation has a positive charge.

An anion has a negative charge.

Critical Thinking Questions:

- Recall the shell model of the atom.
 - How many electrons fill the first shell? _____
 - How many electrons fill the second shell? _____
- How many valence electrons are found in atoms in
 - group 1? _____
 - group 13? _____
 - group 16? _____
- When atoms form ionic compounds, they often fill their outer valence shell (**octet rule**) by either gaining or losing 1, 2, or 3 electrons. Based on your answer to CTQ 2, indicate how many electrons these atoms would either gain or lose to achieve an octet.
 - group 1? _____
 - group 13? _____
 - group 16? _____
- What would be the resulting charge on the ions created in CTQ 3 when electrons are gained or lost?
 - group 1? _____
 - group 13? _____
 - group 16? _____
- Based on your answer to CTQ 4, do metals typically form **anions** or **cations** (circle one)?
- Work with your team and write charges in the empty boxes above each group (skip the transition metals, groups 3-12) at the top of the periodic table in Model 1.
- Do nonmetals typically form **anions** or **cations** (circle one)?

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Richard S. Moog, John J. Farrell



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