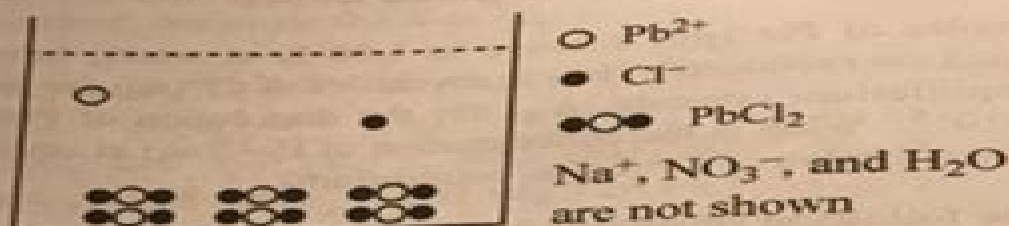


3. Five hundred (500) mL of a  $\text{Pb}(\text{NO}_3)_2$  solution and 500 mL of a  $\text{NaCl}$  solution are added to a large beaker. At equilibrium the result is depicted below:



- Which solution prior to mixture,  $\text{Pb}(\text{NO}_3)_2$  or  $\text{NaCl}$ , had the higher concentration?
- Explain how you know from the diagram that the solution was not prepared by dissolving  $\text{PbCl}_2$  in water.
- Given that each symbol,  $\bigcirc$  and  $\bullet$ , represents  $2.5 \times 10^{-2}$  moles, calculate the  $K_{sp}$  of  $\text{PbCl}_2$  at this temperature.

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**Michael P. Garoutte**

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