

**Prelab Activity:**  
**Chemical Kinetics**

- 1) A student performed experiments 1 and 3 from the table on page 2 of the lab handout. In experiment 1, the reaction mixture took 161 s to turn purple, while in experiment 3, the reaction mixture took 83 s to turn purple.

a) Using Equation 5 in the lab handout, determine the rate of reaction for experiments 1 and 3.

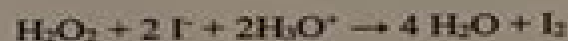
$$\text{Rate} = \frac{1}{2} \frac{\frac{.5\text{mL}}{15\text{mL}} \times .05\text{M}}{\Delta T}$$

$$\text{EXP2} = \frac{1}{2} \frac{\frac{.5\text{mL}}{15\text{mL}} \times .05}{\cancel{161} + \cancel{161} + 83} = 2.00 \times 10^{-4}$$

$$\text{exp1} \quad \frac{1}{2} \frac{\frac{.5\text{mL}}{15\text{mL}} \times .05}{161} = 5.18 \times 10^{-6}$$

- b) Using Equation 2, determine the value of the exponent,  $y$  (the order of the reaction with respect to  $\text{I}^-$ ). Report your value to the nearest whole number. Hint: Remember the total volume of the reaction mixture is 15 mL, so you will need to perform dilution calculations for the  $[\text{I}^-]$  in reactions 1 and 3.

- 2) Consider the following chemical reactions performed in the kinetics lab:



Assuming that you followed the procedure correctly, a color change will occur at approximately what moment?

- When all hydrogen peroxide has been converted to water
- When all of the  $\text{S}_2\text{O}_3^{2-}$  has been used up
- When the  $\text{H}_3\text{O}^+$  has been fully neutralized by the acetate buffer
- When the  $\text{H}_3\text{O}^+$  interacts with the molybdenum catalyst

# Chemfax Kinetics Of A Reaction Prelab Answers

**Ian M. Campbell**



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