

# Pre-Lab Assignment – Lab 11 – Chemical Reactions

Answer: Chadbourne & Davis

1. Match (draw a line) the following chemical reactions to their reaction types:



Combustion

Single-Displacement

Acid-Base

Double-Displacement

Decomposition

2. Briefly describe the actual process in each of the following reaction types:

**Combustion:** A combustion reaction is a reaction in which a hydrocarbon reacts with oxygen gas to form carbon dioxide, water, and heat. Hydrocarbons are very reactive molecules, and so a fairly low-energy combustion reaction will often occur.

**Decomposition:** These are the opposite of decomposition reactions. Thermally.

Thermally, or by heat, the most common decomposition reactions are the production of decomposing through electrolysis. Thermally, thermally, or by heat, the most common decomposition reactions are the production of decomposing through electrolysis.

**Single-Displacement:** A single displacement reaction occurs when one element replaces another in a compound. This type of reaction is the most common reaction. It is the reaction of a metal with a non-metal, or a non-metal with a metal, or a non-metal with a non-metal, or a non-metal with a non-metal.

**Double-Displacement:** In a double displacement reaction, two compounds react to form two new compounds. This type of reaction is the most common reaction. It is the reaction of a metal with a non-metal, or a non-metal with a metal, or a non-metal with a non-metal, or a non-metal with a non-metal.

**Acid-Base:** Acid-base reactions are the most common type of reaction. They are the reaction of an acid with a base, or a base with an acid, or a non-metal with a non-metal, or a non-metal with a non-metal.

# Chemfax Lab Answers Chemical Reactions

**N Colangelo**



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