

KEY

NAME \_\_\_\_\_ PER \_\_\_\_\_

**Lewis Structures and naming practice**

Naming covalent compounds is easy IF you remember six simple prefixes and one suffix. All we do to name covalently bound molecules is tell how many of the different atoms there are in the molecule by adding a prefix, then the whole thing ends with the *-ide* suffix. For example,  $\text{SiO}_2$  is silicon *dioxide*.  $\text{N}_2\text{O}_4$  is *dinitrogen tetroxide*.  $\text{CCl}_4$  is carbon *tetrachloride*. The six you need to know are –

**Di** – to indicate there are two of them

**Tri** – for three

**Tetra** – for four (sometimes we drop the "a" for tetra through deca (10), almost always when the element starts with a vowel– you'll get a feel for it, and I'll never mark you down if you don't)

**Penta** – for five

**Hexa** – for six

**Mono** – for one (only used if there's one of some element and multiples of another.

Then, you put the mono in for the singleton. For example,  $\text{CCl}_3\text{Br}$  is carbon trichloro, monobromide.)

I promise naming's way easier than it seems. Try it with these as you practice Lewis structures. Write the name **or** the formula, and draw the Lewis structure.

<p>Dihydrogen oxide <math>2e^-</math></p> <p><math>\text{H} - \ddot{\text{O}} - \text{H}</math></p> <p><i>You should've known</i></p>	<p>Sulfur trioxide <math>24e^-</math></p> <p><math>\begin{array}{c} \ddot{\text{O}} \\ \parallel \\ \text{:}\ddot{\text{O}} - \text{S} - \ddot{\text{O}}\text{:} \end{array}</math></p>
<p><math>\text{SO}_2</math> Sulfur Dioxide <math>18e^-</math></p> <p><math>\text{:}\ddot{\text{O}} = \ddot{\text{S}} - \ddot{\text{O}}\text{:}</math></p> <p>★</p>	<p><math>\text{NH}_2\text{Cl}</math> <math>5 + 2(1) + 7 = 14e^-</math> Nitrogen dichydrogen chloride</p> <p><math>\begin{array}{c} \text{H} - \ddot{\text{N}} - \text{H} \\   \\ \text{:}\ddot{\text{Cl}}\text{:} \end{array}</math></p>
<p><math>\text{NO}</math> Nitrogen oxide <math>11e^-?</math></p> <p><math>\text{:}\ddot{\text{N}} = \ddot{\text{O}}\text{:}</math></p> <p>★</p>	<p><math>\text{C}_2\text{F}_4</math> <math>2(4) + 4(7) = 36e^-</math></p> <p><math>\begin{array}{c} \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \\ \diagdown \quad \diagup \\ \text{:}\ddot{\text{F}} - \text{C} = \text{C} - \text{F}\text{:} \\ \diagup \quad \diagdown \end{array}</math></p> <p>Dicarbon tetroxide</p>

# Chemfiesta Lewis Structures Answer Key

**Roberts**



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