

pH Calculations – Answers

- 1) Determine the pH of the following solutions:
 - a) A 4.5×10^{-3} M HBr solution.
2.4
 - b) A 3.67×10^{-5} M KOH solution.
9.56
 - c) A solution made by diluting 25 mL of 6.0 M HCl until the final volume of the solution is 1.75 L.
1.1 (the diluted solution is 0.086 M HCl)
 - d) 5 L of an aqueous solution that contains 1.0 grams of HBr and 1.0 grams of nitric acid.
2.2 (the solution has an overall acid concentration of 0.0056 M)
- 2) What are the pOHs for the solutions in problem 1?
 - a) **11.6**
 - b) **4.44**
 - c) **12.9**
 - d) **11.8**
- 3) Explain why even a basic solution contains some H^+ ions.
All aqueous solutions contain H^+ ions from the autoionization of water, $\text{H}_2\text{O} \rightarrow \text{H}^+ + \text{OH}^-$.
- 4) Explain why even an acidic solution contains some OH^- ions.
The same answer from #3 applies here.
- 5) More challenging: What is the pH of a 1.5×10^{-10} M HBr solution?
If you do the calculation to find $-\log(1.5 \times 10^{-10})$, you get an answer of 9.8. However, this is intuitively incorrect – after all, how can a solution that's made of nothing but pure water (pH = 7) with an acid added be basic overall? The answer, it can't. The actual pH of the solution is just about 7, with the main acid source being the H^+ formed from the autoionization of water.
- 6) Does your answer from problem #5 make sense? Explain.
See answer from #5.

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