

Balancing Equations Worksheet – Answers

Note to students: It is acceptable to leave spaces blank when balancing equations – blank spaces are interpreted as containing the number “1”.

- 1) $1 \text{ Na}_3\text{PO}_4 + 3 \text{ KOH} \rightarrow 3 \text{ NaOH} + 1 \text{ K}_3\text{PO}_4$
- 2) $1 \text{ MgF}_2 + 1 \text{ Li}_2\text{CO}_3 \rightarrow 1 \text{ MgCO}_3 + 2 \text{ LiF}$
- 3) $1 \text{ P}_4 + 3 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_3$
- 4) $2 \text{ RbNO}_3 + 1 \text{ BeF}_2 \rightarrow 1 \text{ Be(NO}_3)_2 + 2 \text{ RbF}$
- 5) $2 \text{ AgNO}_3 + 1 \text{ Cu} \rightarrow 1 \text{ Cu(NO}_3)_2 + 2 \text{ Ag}$
- 6) $1 \text{ CF}_4 + 2 \text{ Br}_2 \rightarrow 1 \text{ CBr}_4 + 2 \text{ F}_2$
- 7) $2 \text{ HCN} + 1 \text{ CuSO}_4 \rightarrow 1 \text{ H}_2\text{SO}_4 + 1 \text{ Cu(CN)}_2$
- 8) $1 \text{ GaF}_3 + 3 \text{ Cs} \rightarrow 3 \text{ CsF} + 1 \text{ Ga}$
- 9) $1 \text{ BaS} + 1 \text{ PtF}_2 \rightarrow 1 \text{ BaF}_2 + 1 \text{ PtS}$
- 10) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$
- 11) $2 \text{ NaF} + 1 \text{ Br}_2 \rightarrow 2 \text{ NaBr} + 1 \text{ F}_2$
- 12) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$
- 13) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$
- 14) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$
- 15) $2 \text{ Na}_3\text{PO}_4 + 3 \text{ CaCl}_2 \rightarrow 6 \text{ NaCl} + 1 \text{ Ca}_3(\text{PO}_4)_2$
- 16) $2 \text{ K} + 1 \text{ Cl}_2 \rightarrow 2 \text{ KCl}$
- 17) $2 \text{ Al} + 6 \text{ HCl} \rightarrow 3 \text{ H}_2 + 2 \text{ AlCl}_3$
- 18) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
- 19) $1 \text{ SO}_2 + 2 \text{ Li}_2\text{Se} \rightarrow 1 \text{ SSe}_2 + 2 \text{ Li}_2\text{O}$
- 20) $2 \text{ NH}_3 + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ (NH}_4)_2\text{SO}_4$

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