

## **MIXED GAS LAWS WORKSHEET - SOLUTIONS**

1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ Latm/mol.K})(292 \text{ K})} = 11 \text{ moles of gas}$$

2) If 5.0 moles of  $O_2$  and 3.0 moles of  $N_2$  are placed in a 30.0 L tank at a temperature of  $25^\circ C$ , what will the pressure of the resulting mixture of gases be?  $25^\circ C = 298 \text{ K}$

$$O_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$N_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{tot}} = P_{O_2} + P_{N_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the total moles of gas – what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is  $20.0^\circ C$ . By noon the temperature has risen to  $45.0^\circ C$ . What is the new volume of the balloon?

$$T_1 = 20.0^\circ C = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ C = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ Latm/mol.K})(315 \text{ K})} = 260 \text{ moles of gas}$$

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in  $^\circ C$  of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ Latm/mol.K})} = 296 \text{ K} = 23^\circ C$$

6)  $CaCO_3$  decomposes at  $1200^\circ C$  to form  $CO_2$  gas and  $CaO$ . If 25 L of  $CO_2$  are collected at  $1200^\circ C$ , what will the volume of this gas be after it cools to  $25^\circ C$ ?

$$T_1 = 1200^\circ C = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ C = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at  $20.0^\circ C$  is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is  $-20.0^\circ C$ ?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ C = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ C = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

# Chemfiesta Combined Gas Law Problems Answers

**Patrick Hurley**



## **Chemfiesta Combined Gas Law Problems Answers:**

*Ideal Gas Law 51 Success Secrets - 51 Most Asked Questions on Ideal Gas Law - What You Need to Know* Philip Hicks, 2014-10-07 It's a brand new Ideal gas law world There has never been a Ideal gas law Guide like this It contains 51 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Ideal gas law A quick look inside of some of the subjects covered Atmospheric thermodynamics Overview Thermodynamic instruments Thermodynamic meters Glossary of engineering I Idealization Limits on use Perfect gas Stoichiometry Water vapor Water vapor and dry air density calculations at 0 C Equipartition theorem Perfection Physics and chemistry Glossary of chemistry terms U Fusion energy 1960s Timeline of low temperature technology 19th century Gas Avogadro's law Hot air balloon List of multiple discoveries 17th century Amount of substance Equation of state Overview Explosive Volume of products of explosion Aerodynamics Conservation laws Van der Waals equation Validity Equipartition of energy Gas Physical characteristics Gas meter Flow measurement calculations Mass flow sensor Chamber pressure Importance in Firearm Maintenance Weather forecasting How models create forecasts Timeline of hydrogen technologies 1800s Pressure Pressure of an ideal gas Compressible fluid One Dimensional Flow Diffusion Elementary theory of diffusion coefficient in gases Water vapour Water vapor and Density of airdry air density calculations at 0 C Ideal gas law Numerical weather prediction Computation Gay Lussac's law Pressure temperature law Hydrostatic equilibrium Astrophysics History of thermodynamics Birth of thermodynamics as science and much more

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