

Chemical Arithmetic

- Which one of the following is not an element?
 - Diamond
 - Graphite
 - Silica
 - Ozone
- 10 g CaCO_3 on heating leaves behind a residue weighing 5.6 g. Carbon dioxide released into the atmosphere at STP will be
 - 2.24 L
 - 4.48 L
 - 1.12 L
 - 0.56 L
- 5.6 litres of a gas at N.T.P. are found to have a mass of 11 g. The molecular mass of gas is
 - 22
 - 44
 - 88
 - 32
- The number of atoms of oxygen present in 10.6 g of Na_2CO_3 will be
 - 6.02×10^{22}
 - 12.04×10^{22}
 - 1.806×10^{23}
 - 31.8
- The number of molecules present in a drop of water weighing 0.06 g is approximately
 - 10^{23}
 - 2×10^{23}
 - 3×10^{23}
 - 4×10^{23}
- Which of the following has maximum mass?
 - 0.1 g atom of carbon
 - 0.1 mol of ammonia
 - 6.02×10^{22} molecules of hydrogen gas
 - 1120 cc of carbon dioxide
- A gaseous mixture contains oxygen and nitrogen in the ratio 1 : 4 by weight. Therefore, the ratio of their number of molecules is
 - 1 : 4
 - 1 : 8
 - 7 : 32
 - 3 : 16
- At N.T.P., 5.6 litres of a gas weigh 8 gram. The vapour density of the gas is
 - 32
 - 40
 - 16
 - 8
- On reduction with hydrogen, 3.6 g of an oxide of metal left 3.2 g of the metal. If the atomic weight of the metal is 64, the simplest formula of the oxide would be
 - M_2O_3
 - M_2O
 - MO
 - MO_2
- 280 ml of sulphur at NTP weigh 3.2 g. The molecular formula of the sulphur vapour is
 - S
 - S_2
 - S_8
 - S_4
- The chloride of a metal (M) contains 65.5% of chlorine. 100 ml of the vapour of the chloride of the metal at S.T.P. weigh 0.72 g. The molecular formula of the metal chloride is
 - MCl
 - MCl_2
 - MCl_3
 - MCl_4
- The equivalent weight of a metal is 4.5 and the molecular weight of its chloride is 80. The atomic weight of the metal is
 - 18
 - 9
 - 4.5
 - 36
- The equivalent weight of an element is 4. Its chloride has a vapour density 39.25. Then the valency of the element is
 - 4
 - 3
 - 2
 - 1
- The oxide of an element possesses the formula M_4O_3 . If the equivalent weight of the metal is 9, then the atomic weight of the metal will be
 - 9
 - 18
 - 27
 - none of these
- Approximate atomic weight of an element is 29.89. If its eq. wt. is 8.9, the exact atomic wt. is :
 - 26.89
 - 8.9
 - 17.8
 - 26.7
- 10 g of a piece of marble was put into excess of dilute HCl acid. When the reaction was complete, 1120 cm^3 of CO_2 was obtained at STP. The percentage of CaCO_3 in the marble is
 - 25%
 - 50%
 - 75%
 - 10%

Chemical Arithmetic Answers

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Svarney**



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