

Chemical Equations and Reactions

SECTION 2

SHORT ANSWER Answer the following questions in the space provided.

1. Match the equation type on the left to its representation on the right.

- | | |
|----------------------------------|-----------------------------------|
| <u> c </u> synthesis | (a) $AX + BY \rightarrow AY + BX$ |
| <u> d </u> decomposition | (b) $A + BX \rightarrow AX + B$ |
| <u> b </u> single-displacement | (c) $A + B \rightarrow AX$ |
| <u> a </u> double-displacement | (d) $AX \rightarrow A + X$ |

2. c In the reaction described by the equation $2Al(s) + 3Fe(NO_3)_2(aq) \rightarrow 3Fe(s) + 2Al(NO_3)_3(aq)$, iron has been replaced by

- (a) nitrate. (c) aluminum.
(b) water. (d) nitrogen.

3. a Of the following chemical equations, the only reaction that is both synthesis and combustion is

- (a) $C(s) + O_2(g) \rightarrow CO_2(g)$.
(b) $2C_4H_{10}(l) + 13O_2(g) \rightarrow 8CO_2(g) + 10H_2O(l)$.
(c) $6CO_2(g) + 6H_2O(g) \rightarrow C_6H_{12}O_6(aq) + 6O_2(g)$.
(d) $C_6H_{12}O_6(aq) + 6O_2(g) \rightarrow 6CO_2(aq) + 6H_2O(l)$.

4. b Of the following chemical equations, the only reaction that is both combustion and decomposition is

- (a) $S(s) + O_2(g) \rightarrow SO_2(g)$.
(b) $2C_4H_{10}(l) + 13O_2(g) \rightarrow 8CO_2(g) + 10H_2O(l)$.
(c) $2H_2O_2(l) \rightarrow 2H_2O(l) + O_2(g)$.
(d)

5. Identify the products when the following substances decompose:

- a. **Its separate elements** _____ a binary compound
b. **Metal oxide + water** _____ most metal hydroxides
c. **Metal oxide + carbon dioxide** _____ a metal carbonate
d. **Water + sulfur dioxide** _____ the acid H_2SO_3

6. The complete combustion of a hydrocarbon in excess oxygen yields the products

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Robinson



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