

Worksheet 2: Binary Ionic Compounds (variable charge metals) - metals from groups 3 - 12 and 14 or 4A (Pb and Sn) - charge of anion is constant - use charge of anion to figure out charge of metal

1. Name the following ionic compounds:

a. SnO_2	oxide = O^{2-}	Tin (IV) oxide
b. Mn_2O_7	oxide = O^{2-}	Manganese (VII) oxide
c. FeN	nitride = N^{3-}	Iron (III) nitride
d. Cu_3N_2	nitride = N^{3-}	Copper (II) nitride
e. TiF_3	fluoride = F^{-1}	Titanium (III) fluoride
f. Cu_2S	sulfide = S^{2-}	Copper (I) sulfide
g. Fe_2S_3	sulfide = S^{2-}	Iron (III) sulfide
h. CuBr	bromide = Br^{-1}	Copper (I) bromide
i. Co_3N_2	nitride = N^{3-}	Cobalt (II) nitride
j. CoF_2	fluoride = F^{-1}	Cobalt (II) fluoride

2. Write the formula for the following ionic compounds:

a. Copper (I) nitride	Cu_3N
b. Cobalt (I) fluoride	CoF
c. Titanium (IV) oxide	TiO_2
d. Iron (II) nitride	Fe_3N_2
e. Iron (III) chloride	FeCl_3
f. Copper (II) oxide	CuO
g. Rhodium (II) oxide	RhO
h. Tin (IV) sulfide	SnS_2
i. Manganese (IV) nitride	Mn_3N_4
j. Copper (I) fluoride	CuF
k. Cobalt (II) sulfide	CoS
l. Iron (III) oxide	Fe_3O_2

Chemical Formulas Answer Key

**Bull, Prince Hycy, Patterson, Gerrelyn
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