

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide Cu₂O
2. mercury (I) oxide Hg₂O
3. gold (III) chloride AuCl₃
4. thallium (III) bromide TlBr₃
5. bismuth (V) oxide Bi₂O₅
6. terbium (III) oxide Tb₂O₃
7. uranium (VI) oxide UO₃
8. protactinium (IV) bromide PaBr₄
9. cerium (III) oxide Ce₂O₃
10. arsenic (V) sulphide As₂S₅

11. manganese (III) oxide Mn₂O₃
12. vanadium (II) bromide VBr₂
13. niobium(V) oxide Nb₂O₅
14. titanium (III) oxide Ti₂O₃
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co₃P₂
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr₂
20. copper (II) bromide CuBr₂

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl₄ Tin (IV) Chloride
2. BiBr₅ Bismuth (V) Bromide
3. PoO₂ Polonium (IV) Oxide
4. PbI₂ Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au₂O₃ Gold (III) Oxide
8. FeCl₂ Iron (II) Chloride

9. PdF₄ Palladium (IV) Fluoride
10. Os₂O₃ Osmium (III) Oxide
11. MoBr₂ Molybdenum (II) Bromide
12. VCl₅ Vanadium (V) Chloride
13. Mn₂O₃ Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np₂O₃ Neptunium (III) Oxide
16. V₂O₅ Vanadium (V) Oxide

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Table of Contents Chemical Formulas Practice By Name Answers

1. Understanding the eBook Chemical Formulas Practice By Name Answers
 - The Rise of Digital Reading Chemical Formulas Practice By Name Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Formulas Practice By Name Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Formulas Practice By Name Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Formulas Practice By Name Answers
 - Personalized Recommendations
 - Chemical Formulas Practice By Name Answers User Reviews and Ratings
 - Chemical Formulas Practice By Name Answers and Bestseller Lists
5. Accessing Chemical Formulas Practice By Name Answers Free and Paid eBooks
 - Chemical Formulas Practice By Name Answers Public Domain eBooks
 - Chemical Formulas Practice By Name Answers eBook Subscription Services
 - Chemical Formulas Practice By Name Answers Budget-Friendly Options

6. Navigating Chemical Formulas Practice By Name Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Formulas Practice By Name Answers Compatibility with Devices
 - Chemical Formulas Practice By Name Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Formulas Practice By Name Answers
 - Highlighting and Note-Taking Chemical Formulas Practice By Name Answers
 - Interactive Elements Chemical Formulas Practice By Name Answers
8. Staying Engaged with Chemical Formulas Practice By Name Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Formulas Practice By Name Answers
9. Balancing eBooks and Physical Books Chemical Formulas Practice By Name Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Formulas Practice By Name Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemical Formulas Practice By Name Answers
 - Setting Reading Goals Chemical Formulas Practice By Name Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Formulas Practice By Name Answers
 - Fact-Checking eBook Content of Chemical Formulas Practice By Name Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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