

Chemical Names and Formulas

❖ Key Vocabulary

Monoatomic Ion – Ions formed from a single atom

Binary compound – Compounds composed of two elements

Nomenclature – the “naming system” of binary ionic compounds

Oxyanion – polyatomic ions that contain oxygen

Salt – an ionic compound composed of a cation and the anion from an acid

Chemical Formulas

Subscripts after an elemental symbol represent the number of atoms of the element in the compound

Molecular formulas represent how many atoms of each element are in a compound.
(C_2H_6 , H_2O , $C_6H_{12}O_6$)

Ionic formulas show the simplest ratio, cations to anions, of atoms in the lattice

Parenthesis are used around a polyatomic ion in a chemical formula

Examples $Al_2(SO_4)_3$, $Cu(PO_4)_2$, $Sr(NO_3)_2$

Monoatomic Ions

– Ions formed from a single atom

Group 1 elements may lose one electron = 1+

Group 2 elements may lose two electrons = 2+

Group 13 elements may lose three electrons = 3+

Group 17 elements may gain one electron = 1-

Group 16 elements may gain two electrons = 2-

Group 15 elements may gain three electrons = 3-

Transition metals have various possibilities (no pattern)

Naming monatomic Ions

Monatomic cations are simply called by the element's name

*the stock system is used when there is more than one possibility for an element (more on this later)

Monatomic anions end in IDE such as Fluorine – Fluoride, Oxygen – Oxide
(reference table 1 on page 221)

Binary Ionic Compounds

1. Write the symbols for the ions with the cation first (Al^{3+} O^{2-})

2. Cross over the charges writing the charge of one ion as the subscript for the other ion (Al_2^{3+} O_3^{2-})

3. Multiply subscript and superscript for each ion then find the lowest whole number ratio (Al_2O_3)

Name it!!! Aluminum Oxide

Examples: Zn^{2+} I^- Zn^{2+} S^{2-}

(More on p 223 practice problems)

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Table of Contents Chemical Names And Formulas Chemistry Study Guide

1. Understanding the eBook Chemical Names And Formulas Chemistry Study Guide
 - The Rise of Digital Reading Chemical Names And Formulas Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Names And Formulas Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Names And Formulas Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Names And Formulas Chemistry Study Guide
 - Personalized Recommendations
 - Chemical Names And Formulas Chemistry Study Guide User Reviews and Ratings
 - Chemical Names And Formulas Chemistry Study Guide and Bestseller Lists
5. Accessing Chemical Names And Formulas Chemistry Study Guide Free and Paid eBooks
 - Chemical Names And Formulas Chemistry Study Guide Public Domain eBooks
 - Chemical Names And Formulas Chemistry Study Guide eBook Subscription Services
 - Chemical Names And Formulas Chemistry Study Guide Budget-Friendly Options
6. Navigating Chemical Names And Formulas Chemistry Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Names And Formulas Chemistry Study Guide Compatibility with Devices
 - Chemical Names And Formulas Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Names And Formulas Chemistry Study Guide
 - Highlighting and Note-Taking Chemical Names And Formulas Chemistry Study Guide
 - Interactive Elements Chemical Names And Formulas Chemistry Study Guide

8. Staying Engaged with Chemical Names And Formulas Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Names And Formulas Chemistry Study Guide
9. Balancing eBooks and Physical Books Chemical Names And Formulas Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Names And Formulas Chemistry Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemical Names And Formulas Chemistry Study Guide
 - Setting Reading Goals Chemical Names And Formulas Chemistry Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Names And Formulas Chemistry Study Guide
 - Fact-Checking eBook Content of Chemical Names And Formulas Chemistry Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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