

7**CHEMICAL QUANTITIES****Quiz for CHAPTER 7**

Fill in the word(s) that will make each statement true.

- | | |
|--|--------------|
| 1. Counting, finding mass, and finding volume are all ways of measuring the amount of <u> 1 </u> . | 1. _____ 7.1 |
| 2. The <u> 2 </u> is the mass, in grams, of one mole of atoms of a monatomic element. | 2. _____ 7.1 |
| 3. The term <u> 3 </u> may be used in place of the term gram formula mass to refer to the mass of any element or compound. | 3. _____ 7.2 |
| 4. The representative particle of most elements is the <u> 4 </u> . | 4. _____ 7.2 |
| 5. To convert moles of oxygen gas (O_2) to mass in grams, you can use the conversion factor <u> 5 </u> . | 5. _____ 7.2 |

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- | | |
|---|-----|
| _____ 6. The density of the gas C_2H_6 at STP is 1.34 g/L. | 7.2 |
| _____ 7. The volume of 3 moles of oxygen at STP is 67.2 L. | 7.2 |
| _____ 8. One mole of a gas occupies 22.4 L. | 7.2 |
| _____ 9. To convert a volume of 3.20 L of hydrogen gas to the number of moles, you would use the conversion factor $\frac{22.4 \text{ L}}{1.00 \text{ mol } H_2}$. | 7.2 |
| _____ 10. The density of a gas at STP can be used to calculate the gram formula mass of that gas. | 7.2 |

Solve these problems and show your work in the space provided.

- | | |
|--|-----|
| 11. What mass of gold contains as many atoms as 3.05 g of silver? | 7.2 |
| 12. What is the empirical formula of a compound that is 40.7% carbon, 54.2% oxygen, and 5.1% hydrogen? | 7.3 |

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