

## Chapter 20- Chemical Bonds

- I. Stability of Bonding
  - a. Some properties combine chemically and no longer have the same properties they did before forming a compound.
  - b. A(n) chemical formula is composed of symbols and subscripts indicating the number of atoms of an element in a compound.
  - c. Atoms form compounds when the compound is more stable than the separate atoms.
    - i. Noble gases are more chemically stable than other elements because they have a complete outer energy level.
    - ii. Atoms can lose, gain, or share electrons to get a stable outer energy level.
    - iii. A(n) chemical bond is the force that holds atoms together in a compound.
- II. Types of Bonds
  - a. A(n) ion is charged particle because it has more or fewer electrons than protons.
    - i. When an atom loses an electron, it becomes a positively charged ion; a superscript indicates the charge.
    - ii. When an atom gains an electron, it becomes a negatively charged ion.
  - b. An ionic compound is held together by the ionic bond- the force of attraction between opposite charges of the ions.
    - i. The result of this bond is a neutral compound.
    - ii. The sum of the charges on the ions is zero.
  - c. Molecules are neutral particles formed as a result of sharing electrons.
    - i. A covalent bond is the force of attraction between atoms sharing electrons.
    - ii. Atoms can form double or triple bonds depending on whether they share two or three pairs of electrons.
    - iii. Electrons shared in a molecule are held more closely to the atoms with the larger nucleus.
    - iv. A(n) polar molecule has one end that is slightly negative and one end that is slightly positive although the overall molecule is neutral.
    - v. In a(n) nonpolar molecule electrons are shared equally.
- III. Writing Formulas and Naming Compounds
  - a. Chemists use symbols from the periodic table to write formulas for compounds.
  - b. Binary Compounds- composed of two elements
    - i. Oxidation number- how many electrons an atoms has gained, lost, or shared to become stable
    - ii. Use oxidation numbers and their least common multiples to write formulas.
      - 1. When writing formulas, remember that the compound is neutral.
      - 2. A formula must have the correct number of positive and negative ions so the charges balance.
    - iii. Use the name of the first element, the root name of the second element, and the suffix -*ide* to write the name of the binary ionic compound.
  - c. Polyatomic ion- positively or negatively charged, covalently bonded groups of atoms.
    - i. The compound contains three or more elements.
    - ii. To write names, write the name of the positive ion first; then write the name of the negative ion.
    - iii. To write formulas, use the oxidation numbers, their least common multiple, and put parentheses around the polyatomic ion before adding a subscript.
  - d. Hydrate- compound with water chemically attached to its ions
  - e. Name binary covalent compounds by using prefixes to indicate how many atoms of each element are in the compound.

# Chemical Bonds Chapter 20 Packet Answers

**Steven S. Zumdahl, Paul B. Kelter**



## Chemical Bonds Chapter 20 Packet Answers:

*Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual* T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors' approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors' philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors' approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us "O" Level

*Study Guide - Chemistry Quite Easily Done*, **Organic Chemistry, 13e Student Study Guide and Solutions Manual** T.

W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2022-05-03 Organic Chemistry Student Study Guide and Solutions

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Williams, 2002-01-04 This third edition continues to innovate by providing students with an integrated and modern approach to the subject The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors Also of benefit is the just in time presentation of key content only providing details once they are needed While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons The whole is backed by a range of supplements including a new illustration program a tutorial CD interactive learningware an extensive Web CT component an instructor's resource CD and a solution CD

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