

Chemical Bonds Study Guide

Introduction

Chemical bonds are the forces that hold atoms together in a molecule or compound. Understanding chemical bonding is essential to understanding the behavior of matter in various fields of science. In this study guide, we will cover the different types of chemical bonds and their characteristics.

Types of Chemical Bonds

Covalent Bonds

Covalent bonds occur when atoms share one or more pairs of electrons. The sharing of electrons allows each atom to fill its outermost energy level and become more stable. Covalent bonds can be polar or nonpolar, depending on the electronegativity difference between the atoms.

Ionic Bonds

Ionic bonds occur when one or more electrons are transferred from one atom to another, resulting in the formation of charged ions. The attraction between the opposite charges holds the ions together in a crystal lattice structure.

Metallic Bonds

Metallic bonds occur when metal atoms release their valence electrons to a sea of electrons. The attraction between the atoms and the sea of electrons holds the metal together in a crystal lattice structure.

Bonding Theories

Valence Shell Electron Pair Repulsion (VSEPR)

VSEPR theory predicts the shape of a molecule based on the repulsion between electron pairs in the valence shell. The theory states that electron pairs will arrange themselves to minimize repulsion.

Molecular Orbital Theory

Molecular orbital theory describes the bonding in molecules in terms of the combination of atomic orbitals to form molecular orbitals. It provides a more detailed understanding of the electronic structure of molecules.

Hybridization

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**David E. Sadava, Gordon H.
Orians, Craig Heller, William K. Purves**



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[Fundamentals of Organic Chemistry. Textbook, Study Guide and Solutions Manual](#) T. W. Graham Solomons, 1996-08-15 A realistic approach to the study of

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