

# CHM 116 LAB PROCEDURE

## Chemical Equilibrium: Calculating $K_a$ of a Weak Acid

(~1 – 2 hours)

### Introduction

In this experiment, you will determine the equilibrium constant,  $K_a$ , of a weak, monoprotic acid. The technique used in this experiment is that of **titration**, which is the slow addition of a reactant of known concentration (the **titrant**) to another reactant (the **analyte**) with the purpose of determining an unknown quantity concerning the analyte. The titration you will be performing is one where the titrant is a 0.1 M **sodium hydroxide (NaOH)** solution that will be titrated against **acetic acid (CH<sub>3</sub>COOH)** while the pH of the reaction mixture is monitored throughout the procedure. The **endpoint** of the reaction (the point at which an excess of the **titrant (NaOH)** causes an indicator to change color) will be determined through the observation of a turmeric indicator. The reaction that is taking place in the reaction vessel is:



You will be determining the  $K_a$  (equilibrium constant of a weak acid) and the  $\text{p}K_a$  for the acetic acid using pH and concentration data. Additionally, you will be constructing a **titration curve** for the reaction of a **weak acid (CH<sub>3</sub>COOH)** with a **strong base (NaOH)**, by plotting pH (on the y-axis) against the volume of NaOH added (on the x-axis). The titration curve allows us to determine the equivalence point of the reaction (the point at which the moles of **CH<sub>3</sub>COOH** present and the moles of **NaOH** added are equal). The **equivalence point** is the point on the curve which is half-way up the steep (closest to vertical) portion of the curve. The endpoint and the equivalence point are NOT the same thing, though they should be very close in value.

*Note: The weak acid, **acetic acid**, that you will be using in this lab is an organic acid, and as such the formula is often not written with the same convention as inorganic acids, where the acidic hydrogen is the first element in the formula. In organic acids, the formula is generally written with the acidic hydrogen at the end of the formula, after an oxygen. Thus, the formula for acetic acid is conventionally written as:*



*Where the final **H** is the one which will ionize in water.*

# Chemical Equilibrium Lab Report Answers

**RM Cervero**



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