# Chapter 12: Chemical Equilibrium Quiz

 The equilibrium between carbon dioxide gas and carbonic acid is very important in biology and environmental science.

$$CO2(aq) + H2O(I) \longleftrightarrow H2CO3(aq)$$

Which one of the following is the correct equilibrium constant expression  $(K_d)$  for this reaction?

A) 
$$K_c = \frac{[H_2CO_3]}{[CO_3][H_2O]}$$

D) 
$$K_c = \frac{[co_2]}{[H_2co_3]}$$

B) 
$$K_c = \frac{|co_2|[n_co]}{[n_2co_3]}$$

E) 
$$K_c = \frac{1}{[M_2CO_3]}$$

C) 
$$K_c = \frac{[H_2CO_3]}{[CO_2]}$$

2. The following reactions occur at 500 K. Arrange them in order of increasing tendency to proceed to completion (least → greatest tendency).

1. 
$$2NOC1 \leftarrow \longrightarrow 2NO + Cl_2$$
  $K_p = 1.7 \times 10^{-2}$ 

$$K_P = 1.7 \times 10^{-2}$$

$$K_P = 1.3 \times 10^{-5}$$

$$K_P = 5.9 \times 10^{-4}$$

C) 
$$2 < 3 < 1$$

Consider the two gaseous equilibria:

$$SO_2(g) + 1/2O_2(g) \longleftrightarrow SO_1(g)$$
  
 $2SO_1(g) \longleftrightarrow 2SO_2(g) + O_2(g)$ 

The values of the equilibrium constants K1 and K2 are related by

A) 
$$K_2 = (K_1)^2$$

D) 
$$K_2 = 1/K_1$$

B) 
$$(K_2)^2 = K_1$$

C) 
$$K_2 = 1/(K_1)^2$$

4. When the following reaction is at equilibrium, which of these relationships is always true?

$$2NOCl(g) \leftarrow \longrightarrow 2NO(g) + Cl_2(g)$$

D) 
$$2[NO] = [Cl_2]$$

A) [NO] [Cl<sub>2</sub>] = [NOCI] D) 
$$2[NO] = [Cl_2]$$
  
B) [NO]<sup>2</sup> [Cl<sub>2</sub>] = [NOCI]<sup>2</sup> E) [NO]<sup>2</sup> [Cl<sub>2</sub>] =  $K_e[NOCI]^2$ 

 Calculate K<sub>c</sub> for the reaction 2HI (g) ← → H<sub>2</sub> (g) + I<sub>2</sub> (g) given that the concentrations of each species at equilibrium are as follows: [HI] = 0.85 mol/L, [I<sub>2</sub>] = 0.60 mol/L,  $[H_2] = 0.27 mol/L$ .

# **Chemical Equilibrium Quiz**

**Tao Wei** 

#### **Chemical Equilibrium Quiz:**

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