

### Equilibrium Study Guide



1. Consider the equilibrium above. Which of the following changes will increase the concentration of  $\text{HgI}_4^{2-}$ ?
  - (A) Increasing the concentration of  $\text{OH}^{-}$
  - (B) Adding 6 M  $\text{HNO}_3$
  - (C) Increasing the mass of  $\text{HgO}$  present
  - (D) Increasing the temperature
  - (E) Adding a catalyst
2. In which of the following systems would the number of moles of substances present at equilibrium NOT be shifted by a change in the volume of the system at constant temperature?
  - (A)  $\text{CO(g)} + \text{NO(g)} \leftrightarrow \text{CO}_2(\text{g}) + \frac{1}{2} \text{N}_2(\text{g})$
  - (B)  $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \leftrightarrow 2 \text{NH}_3(\text{g})$
  - (C)  $\text{N}_2(\text{g}) + 2 \text{O}_2(\text{g}) \leftrightarrow 2 \text{NO}_2(\text{g})$
  - (D)  $\text{N}_2\text{O}_4(\text{g}) \leftrightarrow 2 \text{NO}_2(\text{g})$
  - (E)  $\text{NO(g)} + \text{O}_3(\text{g}) \leftrightarrow \text{NO}_2(\text{g}) + \text{O}_2(\text{g})$
3. Which of the following is true for a redox reaction that is thermodynamically favorable at standard conditions?
  - (A)  $\Delta G^{\circ} > 0$ ,  $K_{\text{eq}} > 1$ ,  $E^{\circ}_{\text{cell}} > 0$
  - (B)  $\Delta G^{\circ} > 0$ ,  $K_{\text{eq}} > 1$ ,  $E^{\circ}_{\text{cell}} < 0$
  - (C)  $\Delta G^{\circ} < 0$ ,  $K_{\text{eq}} > 1$ ,  $E^{\circ}_{\text{cell}} > 0$
  - (D)  $\Delta G^{\circ} < 0$ ,  $K_{\text{eq}} < 1$ ,  $E^{\circ}_{\text{cell}} < 0$
  - (E)  $\Delta G^{\circ} < 0$ ,  $K_{\text{eq}} < 1$ ,  $E^{\circ}_{\text{cell}} > 0$
4. For the reaction  $2 \text{A} \leftrightarrow 4 \text{B}$ ,  $K_c = 2.0$ . Determine  $K_c$  for  $4 \text{B} \leftrightarrow 2 \text{A}$ .
  - A) 2.0
  - B) 4.0
  - C) 0.50
  - D) 0.20
  - E) 0.40

# Chemical Equilibrium Study Guide Answer Sheet

**Paul B. Kelter, Steven S. Zumdahl**



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*Fundamentals of Thermodynamics* Mr. Rohit Manglik, 2023-07-23 Explains thermodynamic principles laws and their applications in engineering systems

**Study Guide for Chemical Principles** Thomas Elliott Taylor, 1979

*Study Guide/Selected Solutions Manual* Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

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