

**Chapter
Review****Chemical Reactions****Part A. Vocabulary Review**

Directions: Complete the following sentences using the terms listed below.

- synthesis** **exothermic** **catalyst** **decomposition** **coefficients**
endothermic **single-displacement** **inhibitor** **reactants**
precipitate **chemical** **double-displacement** **products** **balanced**
1. A change in which one or more substances are converted into a new substance is called a(n) **double-displacement** reaction.
 2. If a precipitate, water, or a gas forms when two ionic compounds in solution are combined, the reaction is a(n) **precipitate** reaction.
 3. A substance that prevents certain chemical reactions is a(n) **inhibitor**.
 4. When two or more substances combine in a chemical reaction to form another substance, a(n) **synthesis** reaction has taken place.
 5. The breakdown of a substance into two or more simpler substances is a(n) **decomposition** reaction.
 6. In a chemical equation, the substances that react are called **reactants**.
 7. A substance that speeds up a chemical reaction without itself being permanently changed is a(n) **catalyst**.
 8. A reaction that requires energy in the form of heat is a(n) **endothermic** reaction.
 9. When one element replaces another in a compound, a(n) **single-displacement** reaction has occurred.
 10. If the energy released in a chemical reaction is primarily in the form of heat, that reaction is a(n) **exothermic** reaction.
 11. Numbers in front of the symbols and formulas in an equation are called **coefficients**.
 12. If the number of atoms of each element on the left side of an equation is equal to the number of atoms of each element on the right side of the equation, the equation is **balanced**.
 13. In a chemical equation, the substances that are produced are called **products**.
 14. An insoluble solid that forms when two ionic compounds in solution are combined is a(n) **chemical** reaction.

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