

1869 - Dmitri Mendeleev proposed that properties of elements are in periodic dependence upon their atomic weights.

Mendeleev's Periodic Table:

1. The value of the atomic weight determines the character of the element.
2. There are unknown elements to discover eg., elements similar to Al and Si with atomic weights in range 65-75.
3. The atomic weights of some elements may be changed from knowing the properties of neighbouring elements. Thus the atomic weight of Te must be in range 123-126. It cannot be 128.
4. Some typical properties of an element can be predicted from its atomic weight.

- he predicted the elements: ekaboron (Sc), ekaaluminum (Ga), ekasilicon (Ge) and ekamanganese (Tc)

| | Group I K | Group II K | Group III K | Group IV K | Group V K | Group VI K | Group VII K | Group VIII K |
|---|--------------|---------------|----------------|---------------|--------------|---------------|----------------|-----------------|
| 1 | H=1 | | | | | | | |
| 2 | Li=7 | Be=9 | B=11 | C=12 | N=14 | O=16 | F=19 | |
| 3 | Na=23 | Mg=24 | Al=27 | Si=28 | P=31 | S=32 | Cl=35.5 | |
| 4 | K=39 | Ca=40 | Sc=45 | Ti=48 | V=51 | Cr=52 | Mn=55 | Fe=56 |
| 5 | Rb=85 | Sr=87 | Y=90 | Zr=91 | Nb=93 | Mo=96 | Tc=98 | |
| 6 | Cs=133 | Ba=137 | La=139 | Ce=140 | Pr=141 | Nd=144 | Pm=145 | |
| 7 | Fr=223 | Ra=226 | Ac=227 | Th=232 | Pa=231 | U=238 | Np=237 | |

1870 - Julius Lothar Meyer produced a table of 28 elements (main group of elements) which he listed by their atomic or increasing atomic weights.

- Meyer investigated what was now called transition metals which he listed elements in increasing weight order with elements with the same valence in a column.

Meyer's table in 1870

| | Group I K | Group II K | Group III K | Group IV K | Group V K | Group VI K | Group VII K | Group VIII K |
|---|--------------|---------------|----------------|---------------|--------------|---------------|----------------|-----------------|
| 1 | H=1 | | | | | | | |
| 2 | Li=7 | Be=9 | B=11 | C=12 | N=14 | O=16 | F=19 | |
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| 5 | Rb=85 | Sr=87 | Y=90 | Zr=91 | Nb=93 | Mo=96 | Tc=98 | |
| 6 | Cs=133 | Ba=137 | La=139 | Ce=140 | Pr=141 | Nd=144 | Pm=145 | |
| 7 | Fr=223 | Ra=226 | Ac=227 | Th=232 | Pa=231 | U=238 | Np=237 | |

1870 - Lewis Reeve Gibbs in his paper Synoptical Table of the Chemical Elements, produced a much improved table similar to the modern periodic table of elements

| TABLE OF CHEMICAL ELEMENTS | | | | | | | | | |
|----------------------------|----|----|----|----|----|----|----|----|-----|
| S E C T I O N | | | | | | | | | |
| A | B | C | D | E | F | G | H | I | J |
| 1 | 2 | 3 | 4 | 5 | 6 | 7 | 8 | 9 | 10 |
| 11 | 12 | 13 | 14 | 15 | 16 | 17 | 18 | 19 | 20 |
| 21 | 22 | 23 | 24 | 25 | 26 | 27 | 28 | 29 | 30 |
| 31 | 32 | 33 | 34 | 35 | 36 | 37 | 38 | 39 | 40 |
| 41 | 42 | 43 | 44 | 45 | 46 | 47 | 48 | 49 | 50 |
| 51 | 52 | 53 | 54 | 55 | 56 | 57 | 58 | 59 | 60 |
| 61 | 62 | 63 | 64 | 65 | 66 | 67 | 68 | 69 | 70 |
| 71 | 72 | 73 | 74 | 75 | 76 | 77 | 78 | 79 | 80 |
| 81 | 82 | 83 | 84 | 85 | 86 | 87 | 88 | 89 | 90 |
| 91 | 92 | 93 | 94 | 95 | 96 | 97 | 98 | 99 | 100 |

1913 - Henry Moseley showed a systematic relation between wavelength and atomic number

- Moseley's experiment resulted in arrangement of the periodic table by atomic number rather than atomic weight

1940 - Glenn Seaborg reconfigured the periodic table by placing the lanthanide and actinide series separately from the table

II. The Modern Periodic Table

A. Groupings of Elements

- horizontal rows are called **periods**
- vertical columns are called **groups**
- elements in a group have similar chemical properties

Atomic numbering

- 1-18 (main representative elements)
- 19-10 (transition metal elements)

Atomic numbering: 1 - 10

Group names

- 1A: Alkali Metals
- 2A: Alkaline Earth Metals
- 3A: Boron Group
- 4A: Carbon Group
- 5A: Nitrogen Group
- 6A: Chalcogen Group
- 7A: Halogen Group
- 8A: Noble Gas Group

- elements are arranged according to increasing atomic number
- metals are on the left side of the table zig-zag line
- non-metals are on the right side of the zig-zag line except hydrogen
- metalloids border the zig-zag line except aluminum and gallium
- two rows of elements at the bottom are called inner Transition Elements (lanthanide and actinide series)

Chemical Periodicity Study Guide

Joan Bouillon



Chemical Periodicity Study Guide:

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