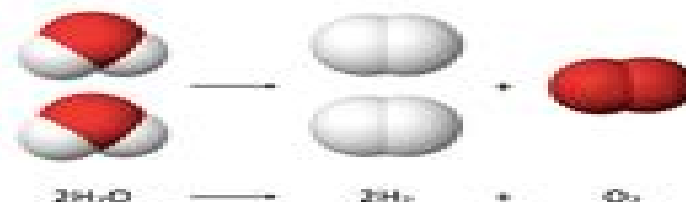


CHEMICAL REACTIONS

Name: _____

In chemistry, we learn about how atoms form bonds with one another in order to satisfy their need for complete electron shells. For example, oxygen's need for 2 additional electrons is filled by the bond formed with two hydrogen atoms, both having one "extra" electron. The bond between these three atoms creates the molecule we know as water. You've been around water your entire life, but have you ever seen a water molecule split back into hydrogen and oxygen?

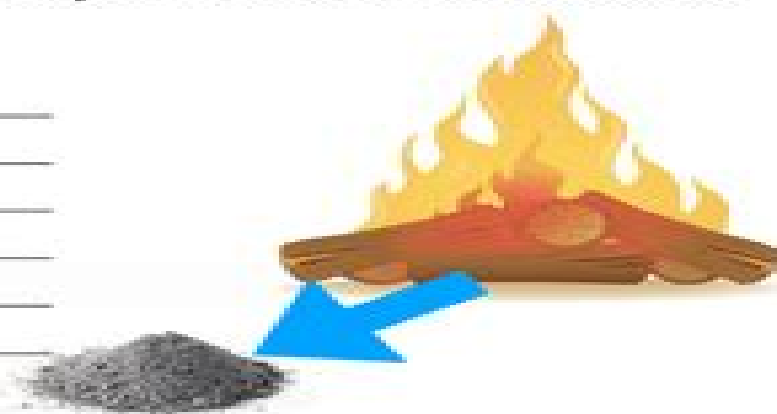
Most likely you haven't! That's because these chemical bonds are not easily broken. Splitting hydrogen and oxygen isn't like picking the pineapple off your pizza. When molecules are broken apart or built up from separate pieces, it's known as a **chemical reaction**. Even if you've never witnessed the splitting of water, chemical reactions happen around you all the time!



When a chemical reaction has occurred, the starting substances (**reactants**) are modified and become something different (**products**). Signs that a chemical reaction have taken place are: temperature increase or decrease, gas production (bubbles), color change, odor, & formation of a precipitate (solids).

Let's try to categorize some examples. Write YES if you think a chemical reaction has occurred, NO if it hasn't.

1. Cutting a piece of paper _____
2. Burning a log _____
3. Rust forming from iron _____
4. Adding sugar to your tea _____
5. Baking a cake _____
6. A banana rotting _____



Chemical reactions involve breaking apart molecules and combining atoms in new ways to form new materials. One thing must always be true of any chemical reaction: The amount of matter that existed at the start of the reaction must equal the amount of matter that exists at the end. This is known as the **Law of Conservation of Mass**. Sound simple? It seems so, but why when you burn a log do you end up with such a small pile of ash? Where did the rest of that matter go? Think about your prediction before you flip the page.

Chemical Reaction Comprehension

H Kauffman



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 - The Rise of Digital Reading Chemical Reaction Comprehension
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Reaction Comprehension
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Reaction Comprehension
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Reaction Comprehension
 - Personalized Recommendations
 - Chemical Reaction Comprehension User Reviews and Ratings
 - Chemical Reaction Comprehension and Bestseller Lists
5. Accessing Chemical Reaction Comprehension Free and Paid eBooks
 - Chemical Reaction Comprehension Public Domain eBooks

- Chemical Reaction Comprehension eBook Subscription Services
- Chemical Reaction Comprehension Budget-Friendly Options
- 6. Navigating Chemical Reaction Comprehension eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Reaction Comprehension Compatibility with Devices
 - Chemical Reaction Comprehension Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Reaction Comprehension
 - Highlighting and Note-Taking Chemical Reaction Comprehension
 - Interactive Elements Chemical Reaction Comprehension
- 8. Staying Engaged with Chemical Reaction Comprehension
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Reaction Comprehension
- 9. Balancing eBooks and Physical Books Chemical Reaction Comprehension
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Reaction Comprehension
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Reaction Comprehension
 - Setting Reading Goals Chemical Reaction Comprehension
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Reaction Comprehension
 - Fact-Checking eBook Content of Chemical Reaction Comprehension
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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