

# Chemical Reactions

## Section 9.1 Reactions and Equations

In your textbook, read about evidence of chemical reactions.

For each statement, write *yes* if evidence of a chemical reaction is present. Write *no* if there is no evidence of a chemical reaction.

- |               |   |
|---------------|---|
| <u>yes</u>    | 1. A tomato smells rotten.  |
| <u>no</u>     | 2. A drinking glass breaks into smaller pieces.                     |
| <u>no</u>     | 3. A piece of ice melts.  |
| <u>yes</u>    | 4. Drain cleaner is mixed with water and the solution becomes warm. |
| <u>yes</u>    | 5. Candle wax burns.  |
| <u>yes/no</u> | 6. Molten candle wax solidifies.                                    |
| <u>yes</u>    | 7. Green leaves turn yellow and red as the seasons change.          |
| <u>yes</u>    | 8. Baking powder produces a gas that makes a cake rise.             |

In your textbook, read about how to represent chemical reactions and how to balance chemical equations.

Use the terms below to complete the passage. Each term may be used once, more than once, or not at all.

arrow	plus sign	(s)	(l)
reactant	product	(g)	(aq)

The fuel for the space shuttle is hydrogen, which burns in oxygen to produce water vapor and energy. In this chemical reaction, hydrogen is a(n) **(9)** reactant, oxygen is a(n) **(10)** reactant, and water vapor is a(n) **(11)** product. In a chemical equation for this reaction, a(n) **(12)** arrow is used to separate hydrogen and oxygen from water vapor and energy. A(n) **(13)** plus sign is used to separate the symbols for hydrogen and oxygen. A(n) **(14)** (g) symbol is used to tell the state of hydrogen in the reaction, a(n) **(15)** (g) symbol is used for the state of oxygen, and a(n) **(16)** (g) symbol is used for the state of water vapor.

# Chemical Reactions Study Guide A Answers

**Robert K. Wismer**



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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*Study Guide for Introductory Chemistry* Iris Stovall, 1990

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**Study Guide and Solutions Manual** Bruce A. Chase, Peter J. Russell, 2005-06 This student resource contains chapter outlines of text material, solutions to all end-of-chapter problems, key terms, suggestions for analytical approaches, problem-solving strategies, and a variety of additional questions for student practice. Also featured are questions that relate to chapter-specific animations and iActivities.

Study Guide with Answers to Selected Problems H. Stephen Stoker, Danny V.

White, 2003-04 This supplement includes for each chapter a brief overview, activities, and practice problems to reinforce skills, and a practice test. The answers section includes answers for all odd-numbered end-of-chapter exercises.

**Study**

**Guide/Selected Solutions Manual** Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter, review of skills, self-tests, and the answers and detailed solutions to all odd-numbered end-of-chapter problems in the text book.

*Study Guide for Chemical Principles [by] Steven S. Zumdahl* Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique

problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples     **Shipman Phys Sci 6e Study Guide** James T. Shipman, Jerry D. Wilson, 1990     **Study Guide** George Karleskint, 1991     **Organic Chemistry, Student Solution Manual and Study Guide** David R. Klein, 2021-03-16

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web worm anatomy a worms body is made up of many segments called annuli the length of a worms body has muscles which contract and relax which enables the worm to move along a surface the annuli are covered in tiny hairs called setae which help the worms movement worms have no lungs so they do not breathe like a human being or

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web 1 find the anterior front end of the earthworm by locating the fleshy bump over its mouth called the prostomium the posterior back end has a small hole where solid waste is expelled called the anus the length of the worm is made up of many tiny segments each separated by a thin wall called a septum 2

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web 1 0 35 annuli what phylum does the earthworm belong to how did this phylum get its name 2 1 01 setae where are setae located and what do they resemble how does the setae enable the earthworm to move 3 1 39 clitellum what is the function of the clitellum

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web dissect the earthworm as described in case of general anatomy after exposing the alimentary canal remove the blood vessels hearts blood glands muscles etc from above and around the alimentary canal now locate the buccal cavity 2nd segment pharynx 3rd to 5th oesophagus 6th 7th gizzard 8th 9th stomach 10th 15th and

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