

## Gas Stoichiometry #4

How many liters of oxygen gas, at 37.0°C and 0.930 atmospheres, can be collected from the complete decomposition of 50.0 grams of potassium chlorate?



$$\frac{50.0 \text{ g } \cancel{\text{KClO}_3}}{122.55 \text{ g } \cancel{\text{KClO}_3}} \times \frac{1 \text{ mol } \cancel{\text{KClO}_3}}{2 \text{ mol } \cancel{\text{KClO}_3}} \times \frac{3 \text{ mol O}_2}{2 \text{ mol } \cancel{\text{KClO}_3}} = 0.612 \text{ mol O}_2$$

$$V = \frac{nRT}{P} = \frac{(0.612 \text{ mol})(0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}})(310 \text{ K})}{0.930 \text{ atm}} = 16.7 \text{ L}$$

# Chemistry B Gas Stoichiometry

**John A. Olmsted, Gregory M.  
Williams, Robert Charles Burk**



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