cf. Slide 9 : text 11x-11 \$

#### I Multiple Choice (20 points)

1. Which of these compounds is a strong electrolyte? - strong base; u. soluble ionic compound

A. H.O.

B. No.

C. CH<sub>3</sub>COOH (acetic acid)

D. C<sub>2</sub>H<sub>6</sub>O (ethanol)

E.KOH

2. Which of these compounds is a weak electrolyte?

A. HCI

B)CH3COOH (acetic acid) C. C.H. Os (glucose)

D. On

E NaCl

3. Which of these compounds is a nonelectrolyte?

A: NaOH

B. HNO:

Oc. HOO; (ethanol) not ionic. Slide 3; slide 9; Sample problem 4.2 (methanol)

slide 7 . Text rp. 112, 113, 116.

weak aid = weak electrolity

E. CH<sub>3</sub>COOH (acetic acid)

Identify the reducing agent in the chemical reaction

 $Cd + NiO_2 + 2H_2O \rightarrow Cd(OH)_2 + Ni(OH)_2$ . cdo -> cd2+ loses e .. oxilized .. reducing agent

A JCd.

B. NiOa C. H.O.

D. CdfOHba

E. Ni(OH):

Sulfur dioxide reacts with chlorine to produce thionyl chloride (used as a drying agent for inorganic halides) and dichlorine oxide (used as a bleach for wood pulp and textiles).

 $SO_2(g) + 2Cl_2(g) \rightarrow SOCl_2(g) + Cl_2O(g)$ 

If 0.400 mol of Cl2 reacts with excess SO2, how many moles of Cl2O are formed?

A. 0.800 mod

B. 0.400 mol. C30,200 mol

D. 0.100 mod. E. 0.0500 mol 0.4 mol Cl2 x 1 mol Cl20 = 0.2 mol Cl20

# **Chemistry 101 Exams And Answer**

Granville Stanley Hall, Edward
Bradford Titchener, Karl M.
Dallenbach, Madison Bentley, Edwin
Garrigues Boring, Margaret Floy
Washburn

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