

A balanced chemical equation is a chemist's recipe.

Chemical Equations					
	$\text{N}_2(\text{g})$	+	$3\text{H}_2(\text{g})$	\longrightarrow	$2\text{NH}_3(\text{g})$
"Microscopic recipe"	1 molecule N_2	+	3 molecules H_2	\longrightarrow	2 molecules NH_3
"Macroscopic recipe"	1 mol N_2	+	3 mol H_2	\longrightarrow	2 mol NH_3

- What would happen if two molecules (moles) of N_2 reacted with three molecules (moles) of H_2 ?

Experimental Conditions		
	Reactants	Products
Before reaction	 2 molecules N_2 3 molecules H_2	0 molecules NH_3

Chemistry 12 3 Limiting Reagent And Percent Yield Answers

David Wilson



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