

Chemistry - 2 Marks questions

Unit - I :-

① write the following

- ② Bond order ③ Heisenberg uncertainty principle ④ Schrodinger wave equation ⑤ significance of ψ and ψ^2 ⑥ HOMO and LUMO

Ans:-

② Bond order:-

Bond order is a measure of the no. of chemical bonds between two atoms in a molecule. It indicates the stability of the bond, with higher bond orders generally corresponding to stronger and shorter bonds.

* Bond order is calculated by the difference b/w the no. of bonding electrons and the no. of anti-bonding electrons divided by 2.

③ Heisenberg uncertainty principle:-

Heisenberg uncertainty principle states that it is impossible to find simultaneously both the position and momentum of the microscopic particles like electron with accuracy.

* Mathematically it can be stated as

$$\Delta x \cdot \Delta p \geq \frac{h}{4\pi} \quad (\because p = mv)$$

$$\Delta x \cdot \Delta v \geq \frac{h}{4\pi m}$$

④ Schrodinger wave equation:-

To provides sense and meaning to the probability approach - Schrodinger derived an equation called Schrodinger wave equation.

To calculate the probability of finding an electron in an atom at various points is the main aim of this equation.

$$\frac{\partial^2 \psi}{\partial x^2} + \frac{\partial^2 \psi}{\partial y^2} + \frac{\partial^2 \psi}{\partial z^2} + \frac{2\pi^2 m}{h^2} (E - PE) \psi = 0$$

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