

CHEM 1202 - Homework # 6

Acids & Bases # 1

Due Thursday, Mar 22, 2022 (by 6 PM)

ANSWER KEY

Check the box to the right if you want your graded homework to be placed out in the audit hall outside Prof. Adams' office. Otherwise, you will have to pick up your homework from Prof. Matter in person.

☐
1. (3 pts) Consider the following weak acids and their pK_a values:Which compound is the strongest acid (use letter)? EThe strongest acid has the lowest or most negative pK_a value.Which compound is the weakest acid (use letter)? DThe weakest acid has the highest pK_a value.Which compound has the strongest conjugate base (use letter)? D

The weakest acid has the strongest conjugate base.

2. (5 pts) What are the pH's for the following solutions?

a) 0.001 M HBr = 3.0

b) 1 M HNO_3 = 0.0

c) 1×10^{-12} M HCl = 7.0

d) 10 M NaOH = 15.0

e) 1×10^{-2} M KOH = 9.0

3. (4 pts) Calculate the pH of a 0.001 M solution of acid that has a pK_a of 5.0. Clearly show all your work and put a box around your answer.

Initial	0.001 M	0 M	0 M
HA(aq)	0.001 - x	x	x
A ⁻ (aq)	0	x	x

$$K_a = \frac{[H^+][A^-]}{[HA]} = 10^{-5.0} = 1 \times 10^{-5}$$

assume $x \ll 0.001$ due to small K_a values

$$x = 1 \times 10^{-4} = [H^+] \quad \boxed{pH = 4.0}$$

Note that x is only one order of magnitude smaller than the concentration (0.001 M), so this approximation just barely works.

4. (4 pts) What is the K_a value of a 1 M solution of carbonic acid (H_2A) that has a pH of 4? Clearly show all your work and put a box around your answer.

Initial	1 M	0 M	0 M
$H_2A(aq)$	1 - x	x	x
$HA(aq)$	0	x	x

BUT, before you go any further, I've given you the pH of the solution in the problem! So you don't have to solve for x , you already know it! So all we have to do is take the anti-log of pH 4 to get the concentration of the H^+ (and A^-).

$[H^+] = \text{anti-log}(-4) = 1 \times 10^{-4}$ M (Substituting this into the equilibrium expression gives us:

$$K_a = \frac{[H^+][A^-]}{[H_2A]} = \frac{1 \times 10^{-4}}{1 - 1 \times 10^{-4}} = 1 \times 10^{-4} \quad \text{so, } \boxed{K_a = 1 \times 10^{-4}} \quad \text{or } pK_a = -\log(K_a) = 4$$

Note that it is fine to drop the x in the $1 - x$ expression because x is much smaller than 1.

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