

Equilibrium and kinetics Notesheet

$K_P > K_c$: at Equilibrium, $[P] > [R]$. Lies to the right, or Favors products.

$K_P = K_c$: at Equilibrium, $[P] = [R]$. Lies to the left, or Favors reactants.

$K_P < K_c$: at Equilibrium, $[R] > [P]$. Lies to the left or Favors reactants.

$$K_c = \frac{[C]^c \cdot [D]^d}{[A]^a \cdot [B]^b} \quad [K_c = K_{eq}]$$

$$K_P = \frac{[C]^c \cdot [D]^d}{[A]^a \cdot [B]^b} \cdot \frac{1}{[P]^p \cdot [Q]^q}$$

$$K_P = K_c (RT)^{\Delta n} \quad (\Delta n = \text{moles of gas, P} - \text{moles of gas, R})$$

$$K_P = K_c \quad \text{if } \Delta n = 0$$

	Stress	Change in K or Q	Results
Equimolar Res ⁿ	Add more reac.	decrease Q	Reaction to the right
	Remove reac.	increase Q	Reaction to the left
	Add more prod.	increase Q	Reaction to the left
	Remove prod.	decrease Q	Reaction to the right
	Increase Temp.	decrease K	Reaction to the left
Submolar Res ⁿ	Decrease Temp.	increase K	Reaction to the right
	Increase Temp.	decrease K	Reaction to the left
	Decrease Temp.	decrease K	Reaction to the left
	Inc. partial p.	shift K	Reaction to fewer gas mole.
	Dec. partial p.	shift K	Reaction to more gas mole.
	Inc. volume	shift K	Reaction to more gas mole.
	Dec. volume	shift K	Reaction to fewer gas mole.
	Adding Catalyst	No change to K or Q	Reach equilibrium sooner

Relating Q , K and ΔG

ΔG	K/Q	Forward Res ⁿ	Characteristics
(-)	$Q < K$	Forward Res ⁿ	Spontaneous, towards products (right)
(+)	$Q > K$	Reverse Res ⁿ	Spontaneous, towards reactants (left)
$\Delta G = 0$	$Q = K$		Reaction at Equilibrium

Identifying acids and bases

A strong acid completely dissociates in water. Weak acids only partially dissociates.

A strong base ionizes completely in water, a weak base only partially ionize.

Bronsted Lowry Definition: Acids \rightarrow donate protons, Bases \rightarrow accept protons

Compounds who act as an acid / base depending on what else is present: amphoteric compounds. Eg. water

$$K_a [\text{acid ionization constant}] = \frac{[A^-][H^+]}{[HA]}$$

$$pK_a = -\log(K_a)$$

$K_a \uparrow$ Acid Strength \uparrow $K_a \uparrow$ pH \downarrow

$pK_a \downarrow$ Acid Strength \uparrow $pK_a \downarrow$ pH \downarrow

$$K_b = \frac{[BH^+][OH^-]}{[B]}$$

$K_b \uparrow$ Base strength \uparrow

$pK_b \downarrow$ Base strength \uparrow

Homogeneous vs Heterogeneous Equilibria

Homogeneous: same phases. Usually gas or aqueous sol.

Heterogeneous: different phases.

K_c consists of only gaseous or aqueous phases of products and reactants.

K vs Q

Q is Reaction Quotient = current conc. of reactants and products.

If $Q > K$ then: proceeds towards the reactants

$Q = K$ then: equilibrium achieved.

$Q < K$ then: proceeds towards products.

Thermodynamics and Equilibrium

At Eq $\Delta G = 0$, $Q = K$

$$\therefore \Delta G = \Delta G^\circ + RT \ln Q$$

$$0 = \Delta G^\circ + RT \ln K$$

$$-RT \ln K = \Delta G^\circ$$

OR

$$K = e^{\left[\frac{-\Delta G^\circ}{RT}\right]}$$

Notesheet

Van't Hoff Equation

$$K = e^{\left[\frac{-\Delta G^\circ}{RT}\right]}$$

$$\ln K = \frac{-\Delta G^\circ}{RT} = \ln K = \frac{-\Delta H^\circ + T\Delta S^\circ}{RT}$$

$$\ln K = \frac{-\Delta H^\circ}{R} \left[\frac{1}{T}\right] + \frac{\Delta S^\circ}{R} \quad \text{slope} = \frac{-\Delta H^\circ}{R}$$

$$\ln \left[\frac{K_2}{K_1}\right] = \frac{-\Delta H^\circ}{R} \left[\frac{1}{T_1} - \frac{1}{T_2}\right] \quad \text{Notesheet}$$

Conjugate acid / base pair

Strong acids have weaker affinity for protons, and have weak conjugate bases.

Weak acids have stronger affinity for protons and have strong conjugate base.

pH and acidity

$$-\log [H^+] = pH$$

$$\downarrow pH = \uparrow \text{Acidity}$$

$$-\log [OH^-] = pOH$$

$$\uparrow pH = \uparrow \text{Basicity}$$

$$pH + pOH = 14$$

$$K_w = [H^+][OH^-]$$

\hookrightarrow ion product constant.

$$\text{At } 25^\circ\text{C (298K)}: K_w = [H^+][OH^-] = 1.0 \times 10^{-14}$$

$$K_a K_b = K_w$$

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Ensheng Dong



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Chemistry 104 Note Taking Guide Answers : Colleen Hoover's "It Ends with Us" This touching tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover masterfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can prevail.

Uncover the Best : Taylor Jenkins Reid's "The Seven Husbands of Evelyn Hugo" This captivating historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reid's absorbing storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery.

Chemistry 104 Note Taking Guide Answers : Delia Owens' "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, entrancing readers with its evocative prose and mesmerizing setting.

These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of compelling stories waiting to be discovered.

The novel begins with Richard Pappen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

The Secret History is a brilliant and gripping novel that will keep you speculating until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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