

Chemistry 103 Worksheet Chapter 2 Isotopes and Ionic Compounds

1. Isotopes

1. Determine the number of protons, neutrons, and electrons in the following isotopes?

(a)  $^{54}_{26}\text{Fe}$

Proton: 26 Neutron: 29 Electron: 26

(b) Carbon-13

Proton: 6 Neutron: 7 Electron: 6

(c) Pb-207

Proton: 82 Neutron: 125 Electron: 82

(d)  $^{79}\text{Br}$

Proton: 35 Neutron: 44 Electron: 35

2. How many protons, neutrons and electrons are present in  $^{17}\text{O}^{2-}$ ?

- ☒ a.  $p=8, n=9, e=10$
- b.  $p=8, n=9, e=8$
- c.  $p=8, n=9, e=6$
- d.  $p=9, n=8, e=11$
- e.  $p=9, n=8, e=9$

3. How many protons, electrons, and neutrons does a  $^{52}\text{Fe}^{2+}$  ion have?

26p 36n 23e<sup>-</sup>

4. X and Y are two species, each consisting of one nucleus and a number of electrons. The two species are found to contain the same number of protons, the same number of neutrons, and different numbers of electrons. Which of the following statements about X and Y is correct?

- a. X and Y are both neutral atoms
- b. X and Y are isotopes of one another

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