

Chemistry Final Exam Study Guide

Chemical Bonding: Ch. 6

Ionic Bonds, Covalent Bonds, Metallic Bonds, Atom Types involved in each bond type, Properties of each bond, Octet Rule, Electron Dot Diagrams, Lewis Dot Structures, Octet Rule Exceptions.

Chemical Formulas & Names: Ch. 7

Naming ions, Binary Ionic Compound Formulas & Names, Ionic compounds with Polyatomic ions-Formulas & Names, Binary Molecular Compound Formulas & Names, Acid Formulas & Names, Oxidation Numbers, % Composition, Empirical Formulas, Molecular Formulas.

Chemical Reactions & Equations: Ch. 8

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Stoichiometry: Ch. 9

Mole Ratios, Mole-Mole Conversions, Mass-Mole Conversions, Mass-Mass Conversions, Volume-Mole Conversions, STP, Particle-Mole Conversions, Limiting Reactant, Percent Yield.

Gases: Ch. 11

Pressure, Pressure Measurement & Units, Dalton's Law of Partial Pressure, Boyle's Law, Charles's Law, Gay-Lussac's Law, Avogadro's Law, Combined Gas Law, Ideal Gas Law, Ideal Gas Constant, Graham's Law.

Solutions, Acids & Bases: Ch. 12, 14, & 15

Solution, Solvent, Solute, Colloid, Suspension, Tyndall Effect, Electrolytes-Strong & Weak, Nonelectrolytes, Concentration, Molarity, Dissociation/Ionization, Hydronium Ion, Acid Properties, Base Properties, Strong Acids & Bases, Weak Acids & Bases, Arrhenius Acid & Base, Bronstad-Lowry Acid & Base, Lewis Acid & Base, mono-, di-, tri- protic acids, Conjugate Acid & Conjugate Base, Neutralization Reaction, Titration, pH, pOH, $[H^+]$, $[OH^-]$.

You will be provided with a card that has monatomic ions and polyatomic ions, a sheet with equations, charts, constants, etc. that you would need to answer certain questions.

You should bring with you 2 or more #2 pencils, a calculator, planner with periodic table, a standard note card (3" x 5") with any information you would like to put onto the card, and your naming compounds/writing chemical formulas flowcharts.

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Lyn Baker



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