

Practice Final

Prof. Brindley

12.14.2022

10. What is the molar concentration of chloride ions in a solution prepared by mixing 100.0 mL of 2.0 M KCl with 50.0 mL of a 1.50 M  $\text{CaCl}_2$  solution?

$$\begin{aligned}
 0.1 \text{ L} \cdot \frac{2.0 \text{ mol KCl}}{1 \text{ L}} &= 0.2 \text{ mol Cl}^- \\
 0.05 \text{ L} \cdot \frac{1.5 \text{ mol CaCl}_2}{1 \text{ L}} \cdot \frac{2 \text{ mol Cl}^-}{2 \text{ mol CaCl}_2} &= 0.15 \text{ mol Cl}^- \\
 \hline
 0.2 \text{ mol Cl}^- + 0.15 \text{ mol Cl}^- &= 0.35 \text{ mol Cl}^- \\
 \hline
 \frac{0.35 \text{ mol Cl}^-}{0.15 \text{ L}} &= 2.33 \text{ M Cl}^-
 \end{aligned}$$

11. To what volume should you dilute 133 mL of an 7.90 M  $\text{CuCl}_2$  solution so that 51.5 mL of the diluted solution contains 4.49 g  $\text{CuCl}_2$ ?

$$\begin{aligned}
 4.49 \text{ g CuCl}_2 \cdot \frac{1 \text{ mol}}{134.45 \text{ g}} &= 0.0334 \text{ mol CuCl}_2 \\
 0.0334 \text{ mol CuCl}_2 &= 0.0515 \text{ L} \cdot 7.90 \text{ M} = 0.648 \text{ M} \\
 M_1 V_1 &= M_2 V_2 \rightarrow 7.90 \text{ M} \cdot 0.133 \text{ L} = 0.648 \text{ M} \cdot V_2 \\
 V_2 &= \frac{7.9 \cdot 0.133}{0.648} = 1620 \text{ mL} \\
 \text{Dilute to } 1620 \text{ mL} \\
 \text{Add } 1620 - 133 \text{ mL of water} &= 1487 \text{ mL}
 \end{aligned}$$

12. In lab you performed a titration of acetic acid with sodium hydroxide. 83 mL of sodium hydroxide at a concentration of 0.015M were needed to neutralize 58.2 mL of an acetic acid solution. What is the concentration of the acetic acid solution?

# Chemistry 1230 Practice Final

**TD Snyder**



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*Principles and Practices of Winemaking* Roger B. Boulton, 1996 This essential text and reference offers a complete guide to winemaking The authors all well known experts in their field concentrate on the process of wine production stressing the chemistry biochemistry microbiology and underlying science of enology They present in depth discussion of every aspect of the wine production process from the selection of grapes and preparation of the must and the juice through aging bottling and storage of finished wines Novices and experienced winemakers alike will find this clearly written and expertly crafted book an indispensable source of practical instruction and information **Chemical Engineering Catalog** , 1919

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Eugene Franz Roeber, Howard Coon Parmelee, 1919 *Industrial Chemist and Chemical Manufacturer* , 1927 **The Industrial Chemist and Chemical Manufacturer** , 1927 **Federal Practice and Procedure** Charles Alan Wright, 1969

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