

Section Review

Objectives

1. Explain the formation of ionic compounds.
2. Describe the properties of ionic compounds.

Vocabulary

- | | |
|-------------------|-------------------|
| a. ionic compound | b. cation |
| c. anion | d. ionic compound |
| e. ionic crystal | |

Part A Completion

Write the complete name of each compound, and describe each compound and how it is composed. Write down each compound's chemical formula and its physical properties.

1. Sodium and chlorine atoms are needed for sodium chloride. The name of sodium chloride is NaCl. Sodium chloride is a white crystalline solid. Sodium chloride is used for table salt. Sodium chloride is a crystalline solid. Sodium chloride is a crystalline solid.

2. Magnesium and oxygen atoms are needed for magnesium oxide. The name of magnesium oxide is MgO. Magnesium oxide is a white crystalline solid. Magnesium oxide is used for refractory materials. Magnesium oxide is a crystalline solid. Magnesium oxide is a crystalline solid.

Part B True-False

Write true or false next to each statement. If the statement is false, correct the statement.

1. True The ionic bond is the force of attraction between the positive and negative ions in an ionic compound.

3. The ionic bond is the force of attraction between the positive and negative ions in an ionic compound.
4. The ionic bond is the force of attraction between the positive and negative ions in an ionic compound.
5. The ionic bond is the force of attraction between the positive and negative ions in an ionic compound.
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12. The ionic bond is the force of attraction between the positive and negative ions in an ionic compound.

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