Equilibrium Study Guide

$$HgO(s) + 4 \Gamma(aq) + H_2O(l) \rightarrow HgI_4^{2-}(aq) + 2 OH^{-}(aq)$$
 $\Delta H < 0$

- Consider the equilibrium above. Which of the following changes will increase the concentration of Hgl₄⁻?
 - (A) Increasing the concentration of OH⁻
 - (B) Adding 6 M HNO₃
 - (C) Increasing the mass of HgO present
 - (D)Increasing the temperature
 - (E) Adding a catalyst
- 2. In which of the following systems would the number of moles of substances present at equilibrium NOT be shifted by a change in the volume of the system at constant temperature?
 - $(A) CO(g) + NO(g) \leftrightarrow CO_2(g) + \frac{1}{2} N_2(g)$
 - (B) $N_2(g) + 3 H_2(g) \leftrightarrow 2 NH_3(g)$
 - (C) $N_2(g) + 2 O_2(g) \leftrightarrow 2 NO_2(g)$
 - $(D)N_2O_4(g) \leftrightarrow 2NO_2(g)$
 - (E) $NO(g) + O_3(g) \leftrightarrow NO_2(g) + O_2(g)$
- 3. Which of the following is true for a redox reaction that is thermodynamically favorable at standard conditions?
 - (A) $\Delta G^* > 0$, $K_{eq} > 1$, $E^*_{cell} > 0$
 - (B) $\Delta G^* > 0$, $K_{eq} > 1$, $E^*_{cell} < 0$
 - (C) $\Delta G^* < 0$, $K_{eq} > 1$, $E^*_{cell} > 0$
 - (D) $\Delta G^* < 0$, $K_{eq} < 1$, $E^*_{cell} < 0$
 - (E) $\Delta G^{\circ} < 0$, $K_{eq} < 1$, $E^{\circ}_{cell} > 0$
- 4. For the reaction 2 A ↔ 4 B, K_c = 2.0. Determine K_c for 4 B ↔ 2 A.
 - A) 2.0
 - B) 4.0
 - C) 0.50
 - D) 0.20
 - E) 0.40

Chapter 1chemical Equilibrium Study Guide Answers

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