

CHEMISTRY 10TH – Short / Detailed Question Answers

→ CHEMICAL EQUILIBRIUM

CHAPTER# 01

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Q.1: What do you know about chemical reactions?**Ans: Chemical Reactions:**

Many physical and chemical changes are occurring in our surroundings which may be due to chemical reactions. In these reactions, reactants are converted into one or more products.

A chemical reaction is a chemical change which involves reactants and products.

For example : The formation of water from hydrogen gas and oxygen gas, the decomposition reaction of sodium bicarbonate into sodium carbonate, water, and carbon dioxide etc. A chemical reaction contains two quantities; reactant and product which are separated by an arrow.



The reactions may be reversible (condensation, evaporation, freezing, melting) or irreversible (combustion, rusting).

Q.2: Define chemical equilibrium with example**Ans: Chemical Equilibrium:**

The reversible chemical reaction never goes to completion because the product reacts and reproduces reactants again and takes place in forward and reverse directions. This state in which forward reaction rate and reverse reaction rate are equal is known as equilibrium. We know that equilibrium means a balance. Equilibrium exists in many ways in our surroundings.

It is a state where the rate of the forward reaction is equal to the rate of backward reaction OR it is also defined as "A state where the concentration of reactant and product remains constant."

Example: $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons \text{HI}_{(g)}$

Chapter 1review Chemical Equilibrium Mixed Answers

Peipei Pang



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Applied Mechanics Reviews ,1996 Modern Chemistry Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001

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Jespersen,Pamela Kerrigan,2023-07-04 A guide to taking the Advanced Placement exam in chemistry featuring a review of major chemistry concepts practice and diagnostic tests test taking strategies an overview of the test and practice problems

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spectrophotometer are included Updated instructions for Excel spreadsheets to Excel 2007 **Scientific and Technical Aerospace Reports** ,1986 *Biopolymer Mixtures* Stephen E. Harding,Sandra E. Hill,John Richard Mitchell,1995 Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl,Paul B. Kelter,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples General, Organic, and Biochemistry William Henry Brown,Elizabeth P. Rogers,1980 *New Technical Books* New York Public Library,1978 Chemistry John W. Moore,Conrad L. Stanitski,Peter C. Jurs,2002

CHEMISTRY THE MOLECULAR SCIENCE is intended to help students develop a broad overview of chemistry and chemical reactions an understanding of the most important concepts and models that chemists and those in chemistry related fields use an appreciation of the many ways chemistry impacts our daily lives the ability to apply the facts concepts and models of chemistry appropriately to new situations in chemistry other sciences and engineering and to other disciplines **Annual Review of Physical Chemistry** ,1968 **Bulletin of Chemical Thermodynamics** ,1984 **Medical Council** ,1909 **The Electrical Review** ,1906 **The Chamber of Commerce Journal** ,1921 **Annual Review of Physical Chemistry** Gerhard Krohn Rollefson,1968 Provides abstracts and review articles on topics in physical chemistry *Iron & Coal Trades Review* ,1926 Vol 115 includes Diamond jubilee issue 1867 1927 Physical Chemistry John S. Winn,1995 This text presents physical chemistry as a coherent whole rather than a set of disjointed topics and shows how the subject relates to the rest of chemistry and physics It emphasizes physical models as well as mathematical techniques along with both rigorous and approximate order of magnitude problem solving Designed to progress beyond a numerical answer problems expose the physical significance of the situation and teach students how to pose a problem in the first place In addition modern molecular concepts currently unanswered problems in research experimental techniques and new directions in the field are introduced wherever appropriate An orderly progression of thermodynamics carefully builds students knowledge without covering too much too early on Chemical reaction thermodynamics is covered in Chapter 7 after the culmination of thermodynamics with advanced material in Chapter 10

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