

$$dG = -SdT + VdP + \sum_{i=1}^c \mu_i dN_i$$

The Gibbs Equation

$$0 = \sum_{i=1}^c \mu_i dN_i$$

When a system is in Equilibrium, the change in Gibb's energy is 0. In our example, neither the volume nor the temperature of the system change.

$$0 = \sum_{i=1}^c \mu_i dN_i = \mu_1 + \mu_2 + \dots + \mu_c$$

An expansion of the total change in chemical potential for each component with respect to the change in phase.

$$0 = \mu_1 = \mu_1^{(1)} + \mu_1^{(2)} + \dots + \mu_1^{(N_p)}$$

Because our system consists of only one component, the total chemical potential is only that contained in the component (1).

$$0 = \mu_1^{(1)} + \mu_1^{(2)}$$

Because we only have two phases, the summation of the chemical potential of those two phases is equal to 0.

$$\mu_1^{(1)} = -\mu_1^{(2)}$$

This allows us to see that at Equilibrium, the chemical potential of the liquid phase and the vapor phase are equal. The signs are opposite because on change in phase would release the Gibbs Energy that is taken in by the other during a change in phase.

Chapter 1section 4 Phase Equilibrium Answers

Kathleen E. Murphy



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Thermodynamics Problem Solving in Physical Chemistry Kathleen E. Murphy, 2020-03-23 Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation create appropriate solutions and gain confidence through practice solving physical chemistry problems The workbook includes six major sections with 20 30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area Key Features Provides instructor access to a visual map depicting how all equations used in thermodynamics are connected and how they are derived from the three major energy laws Acts as a guide in deriving the correct solution to a problem Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry Can be used as a stand alone product for review of Thermodynamics questions for major tests

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Fundamentals of Physics, Part 4, Chapters 34-38

David Halliday, Robert Resnick, Jearl Walker, 1996-08-16 Part 3 of the fifth edition of this introduction to physics. This text addresses the issue of building bridges of reason so that students may move from qualitative understanding of any given physics concept to making decisions about how to solve a problem involving that concept.

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MCAT Comprehensive Review, 2004 Kaplan, Inc., 2003 The bestselling MCAT prep guide is a must for pre

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