

Chemistry – Dr. Hazlett
Acids/Bases - Chapter 19
Study Guide

I. Introduction

- A. Acids (A) and Bases (B) are in solutions (solns)
 - 1. Solns are homogeneous mixtures of 2+ substances
 - a. Solute and solvent
 - 2. Electrolytes and non-electrolytes
 - a. Aqueous (aq) solns of ionic compounds conduct electricity – i.e. flow of electrons, are electrolytes
 - b. Pure (distilled) water not a good conductor – no ionization and therefore, no free electrons
 - c. A's ionize in water – thus are electrolytes
 - (1) Ionization varies up to 100%
 - (2) Weaker acids ionize less than 100%
 - 3. Ionization
 - a. Aqueous soln (water) → end up with H⁺ and OH⁻ ions
 - b. The amounts of these ions in soln determine whether A or B
 - c. Water is universal solvent
 - (1) Undergoes self-ionization
 - (a) $\text{H}_2\text{O} \rightarrow \text{H}^+ + \text{OH}^-$
 - (b) $\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{OH}^-$

II. Acids and Bases in General

A. Characteristics

- 1. Acids - Physical:
 - a. Taste sour
 - b. React with some metals (Al, Mg, Zn) to produce H₂ (g)
 - c. React with carbonates to produce CO₂ (g)
 - d. React with B to form salts and water
 - e. Change indicators to differing colors – most common is litmus blue to red
 - f. Yields H⁺ ion or forms H₃O⁺
- 2. Bases – Physical:
 - a. Taste bitter
 - b. Feel soapy/slippery
 - c. Changes red litmus paper to blue
 - d. Reacts with A's to form salts and water (neutralization)
 - e. Reacts with some metals to produce H₂ (g)
 - f. Yields the OH⁻ ion
 - g. Alkalines

B. History and Definitions (Source:

<http://library.thinkquest.org/C006669/data/Chem/acidbases/development>).

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PT Brinkman



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