

20**ACIDS AND BASES**
PRACTICE PROBLEMS

In your notebook, solve the following problems.

SECTION 20.1 DESCRIBING ACIDS AND BASES

1. Name the following compounds as acids.
 - a. HNO_2
 - b. H_2SO_4
 - c. HF
 - d. H_2CO_3
2. Write the formulas for the following bases.
 - a. calcium hydroxide
 - b. ammonium hydroxide
 - c. aluminum hydroxide
 - d. lithium hydroxide
3. Describe some distinctive properties of acids.
4. Describe some distinctive properties of bases.

SECTION 20.2 HYDROGEN IONS AND ACIDITY

1. A solution has a hydrogen ion concentration of $1 \times 10^{-6}\text{M}$. What is its pH?
2. What is the pH of a solution if the $[\text{H}^+] = 7.2 \times 10^{-9}\text{M}$?
3. What is the pOH of a solution if the $[\text{OH}^-] = 3.5 \times 10^{-3}\text{M}$?
4. What is the pOH of a solution that has a pH of 3.4?
5. Classify each solution as acidic, basic, or neutral.
 - a. $[\text{H}^+] = 2.5 \times 10^{-9}\text{M}$
 - b. $\text{pOH} = 12.0$
 - c. $[\text{OH}^-] = 9.8 \times 10^{-11}\text{M}$
 - d. $[\text{H}^+] = 1 \times 10^{-7}\text{M}$
 - e. $\text{pH} = 0.8$
6. Calculate the pH of each solution.
 - a. $[\text{H}^+] = 1 \times 10^{-5}\text{M}$
 - b. $[\text{H}^+] = 4.4 \times 10^{-11}\text{M}$
 - c. $[\text{OH}^-] = 2.2 \times 10^{-7}\text{M}$
 - d. $\text{pOH} = 1.4$
7. Classify the solutions in problem 6 as acidic or basic.
8. Why is there a minus sign in the definition of pH?
9. A solution has a pOH of 12.4. What is the pH of this solution?
10. What is the pH of a solution with $[\text{H}^+] = 1 \times 10^{-9}\text{M}$?

SECTION 20.3 ACID-BASE THEORIES

1. Identify the hydrogen ion donor(s) and hydrogen ion acceptor(s) for ionization of H_2SO_4 in water. Label the conjugate acid-base pairs.
2. Identify all of the ions that may be formed when H_3PO_4 ionizes in water.

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