

Chapter 20 Worksheets: Redox ANSWERS

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

- $2\text{Sr} + \text{O}_2 \longrightarrow 2\text{SrO}$ $\text{Sr}^0 \text{ to } \text{Sr}^{2+}$; oxidized/reducing agent $\text{O}^0 \text{ to } \text{O}^{2-}$; reduced/ox. ag.
- $2\text{Li} + \text{S} \longrightarrow \text{Li}_2\text{S}$ $\text{Li}^0 \text{ to } \text{Li}^{+}$; oxidized/ox. ag. $\text{S}^0 \text{ to } \text{S}^{2-}$; reduced/ox. ag.
- $2\text{Cu} + \text{Br}_2 \longrightarrow 2\text{CuBr}$ $\text{Cu}^0 \text{ to } \text{Cu}^{+}$; oxidized/ox. ag. $\text{Br}^0 \text{ to } \text{Br}^{-}$; reduced/ox. ag.
- $3\text{Mg} + \text{N}_2 \longrightarrow \text{Mg}_3\text{N}_2$ $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$; oxidized/ox. ag. $\text{N}^0 \text{ to } \text{N}^{3-}$; reduced/ox. ag.
- $4\text{Fe} + 3\text{O}_2 \longrightarrow 2\text{Fe}_2\text{O}_3$ $\text{Fe}^0 \text{ to } \text{Fe}^{3+}$; oxidized/ox. ag. $\text{O}^0 \text{ to } \text{O}^{2-}$; reduced/ox. ag.
- $\text{Cl}_2 + 2\text{NaBr} \longrightarrow 2\text{NaCl} + \text{Br}_2$ $\text{Cl}^0 \text{ to } \text{Cl}^{-}$; reduced/ox. ag. $\text{Br}^{-} \text{ to } \text{Br}^0$; oxidized/ox. ag.
- $\text{Si} + 2\text{F}_2 \longrightarrow \text{SiF}_4$ $\text{Si}^0 \text{ to } \text{Si}^{4+}$; oxidized/ox. ag. $\text{F}^0 \text{ to } \text{F}^{-}$; reduced/ox. ag.
- $\text{Mg} + 2\text{HCl} \longrightarrow \text{MgCl}_2 + \text{H}_2$ $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$; oxidized/ox. ag. $\text{H}^{+} \text{ to } \text{H}^0$; reduced/a.
- $2\text{Na} + 2\text{H}_2\text{O} \longrightarrow 2\text{NaOH} + \text{H}_2$ $\text{Na}^0 \text{ to } \text{Na}^{+}$; oxidized/a. $\text{H}^{+} \text{ to } \text{H}^0$; reduced/a.

11. Give the oxidation number of each kind of atom or ion.

- | | | | | | | |
|------------|-------|--------------------|---------------------|---------------------|------------|-------------|
| a. sulfate | b. Sn | c. S^{2-} | d. Fe^{3+} | e. Sn^{4+} | f. nitrate | g. ammonium |
| 2- | 0 | 2- | 3+ | 4+ | 1- | 1+ |

12. Calculate the oxidation number of chromium in each of the following.

- | | | | | |
|----------------------------|---------------------------------------|--------------------|-------------|---------------|
| a. Cr_2O_3 | b. $\text{Na}_2\text{Cr}_2\text{O}_7$ | c. CrSO_4 | d. chromate | e. dichromate |
| 3+ | 6+ | 2+ | 7+ | 6+ |

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)

- $\text{C} + \text{H}_2\text{SO}_4 \longrightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$ $\text{C}^0 \text{ to } \text{C}^{4+}$; oxidized $\text{S}^{6+} \text{ to } \text{S}^{4+}$; reduced
- $\text{HNO}_3 + \text{HI} \longrightarrow \text{NO} + \text{I}_2 + \text{H}_2\text{O}$ $\text{N}^{5+} \text{ to } \text{N}^{2+}$; reduced $\text{I}^{-} \text{ to } \text{I}^0$; oxidized
- $\text{KMnO}_4 + \text{HCl} \longrightarrow \text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O} + \text{KCl}$ $\text{Mn}^{7+} \text{ to } \text{Mn}^{2+}$; reduced $\text{Cl}^{-} \text{ to } \text{Cl}^0$; oxidized
- $\text{Sn} + \text{HNO}_3 \longrightarrow \text{SnO}_2 + \text{NO} + \text{H}_2\text{O}$ $\text{Sn}^0 \text{ to } \text{Sn}^{4+}$; oxidized $\text{N}^{5+} \text{ to } \text{N}^{2+}$; red.

14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

- | | |
|-------------------------------|--------------------------|
| a. oxidizing agent: sulfur | reducing agent: carbon |
| b. oxidizing agent: nitrogen | reducing agent: iodine |
| c. oxidizing agent: manganese | reducing agent: chlorine |
| d. oxidizing agent: nitrogen | reducing agent: antimony |

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