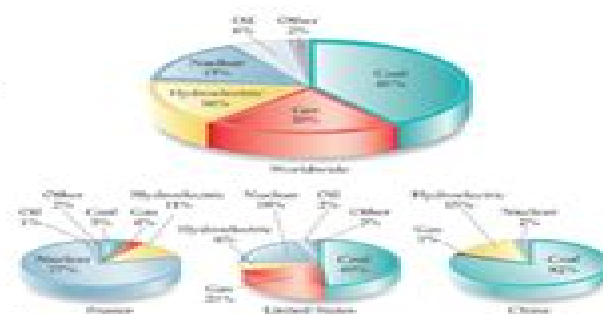


CHEM 1412 Chapter 21 Notes: Nuclear

21.1 Nuclear Structure and Stability

The Nucleus

- Nuclear chemistry is the study of reactions that involve changes in nuclear structure.
- When referring to a single type of nucleus, we often use the term **nuclide** and identify it by the notation A_ZX . Remember that the nucleus is composed of the two **nucleons**, protons and neutrons.
 - The number of protons (Z) is the **atomic number**.
 - The number of neutrons is N .
 - The number of protons and neutrons (A) ($A=Z+N$) together is effectively the **mass number** of the atom.
- Isotopes**
 - Not all atoms of the same element have the same mass, due to different numbers of neutrons in those atoms.
 - There are, for example, three naturally occurring isotopes of uranium:
 - Uranium-234
 - Uranium-235
 - Uranium-238
 - It is not uncommon for some nuclides of an element to be unstable



Mass number (number of protons plus neutrons)

Atomic number (number of protons or electrons)

${}^{12}_6\text{C}$ — Symbol of element

Neutron-Proton Ratios

- Any element with more than one proton (i.e., anything but hydrogen) will have repulsions between the protons in the nucleus.
- Positively charged protons that repel one another at such short distances requires very strong attractive forces
- A **strong nuclear force** helps keep the nucleus from flying apart.
- Neutrons play a key role stabilizing the nucleus.
- Therefore, the ratio of neutrons to protons is an important factor.
- For smaller nuclei ($Z \leq 20$), stable nuclei have a neutron-to-proton ratio close to 1:1.
- As nuclei get larger, it takes a larger number of neutrons to stabilize the nucleus.

$$E=mc^2$$

For ${}^4_2\text{He}$

$$(2 \times 1.0073 \text{ amu}) + (2 \times 1.0087 \text{ amu}) + (2 \times 0.00055 \text{ amu}) = 4.0331 \text{ amu}$$

protons neutrons electrons

- But the mass spectrometric shows 4.0026 amu
- Mass defect** is the difference between the calculated and experimentally measured masses.
 - the mass defect indicates a "loss" for helium in mass of $4.0331 \text{ amu} - 4.0026 \text{ amu} = 0.0305 \text{ amu}$.
- The loss in mass accompanying the formation of an atom from protons, neutrons, and electrons is due to the conversion of that mass into energy that is evolved as the atom forms.
- The **nuclear binding energy** is the energy produced when the atoms' nucleons are bound together
 - this is also the energy needed to break a nucleus into its constituent protons and neutrons.
- The conversion between mass and energy equation by Einstein: $E=\Delta mc^2$
 - where E is energy in Joules, Δm is the mass defect in kg ($6.022 \times 10^{23} \text{ amu} = 1 \text{ g}$ and $1000 \text{ g} = 1 \text{ kg}$), and c is the speed of light in a vacuum, $2.998 \times 10^8 \text{ m/s}$.

Learning Check

The nuclear mass of ${}^4_2\text{He}$ is 4.0026 amu. Calculate the binding energy per nucleon (J/nucleon).

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DJ Losen



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