

I. Chemical Changes

a. Describing chemical reactions- change of one or more substances converted into new substances

1. Reactants are substances that combine or change.
2. New substances that are produced are called products.

b. Conservation of Mass—a law which states that, in a chemical reaction, matter is not created or destroyed; it stays the same.

1. Antoine Lavoisier—experimented with mercury (II) oxide and heat
2. Found mass of products (liquid mercury and oxygen gas) equaled mass of reactants

c. Writing equations—a chemical equation uses chemical formulas and symbols to describe a chemical reaction and the product(s) it produces.

1. Chemical formula expresses the relationship between elements in the compounds and molecules they make up
2. Coefficients—numbers which represent the number of units of each substance in a reaction
3. Knowing coefficients of chemical reactions allows chemists to use the correct amounts of reactants to predict the amounts of products.
4. Subscripts—numbers which represent the number of atoms in a molecule of a particular element.
5. Symbols used to show state of reactants: (s) solid, (aq) aqueous, (g) gas, (l) liquid.

d. Metals react with atmosphere in different ways.

II. Chemical Equations

a. Checking for balance—law of conservation of mass requirement

1. A balanced chemical reaction—both sides of equation have same number of atoms of each element
2. Choosing coefficients— becomes easier with practice; trial and error at first

b. Writing balanced chemical equations—a four-step process

1. Describe the reaction in words.
2. Write the equation using formulas and symbols.
3. Check for balance.
4. Add coefficients where needed for balance.

Chapter 21 Resource Chemical Reactions Answer Key

Michael J. Padilla



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