

Exam:

Name: _____

MULTIPLE CHOICE Choose the one alternative that best completes the statement or answers the question.

- 1) The total concentration of ions in a 0.250 M solution of HCl is _____ 11 _____
A) 0.250 M
B) 0.500 M
C) 0.125 M
D) 0.750 M
E) essentially zero.
- 2) A weak electrolyte exists predominantly as _____ in solution. 21 _____
A) electrons
B) molecules
C) atoms
D) ions
E) an isotope
- 3) Which of the following are strong electrolytes? 31 _____
HCl
HC₂H₃O₂
NH₃
KCl
A) HCl, HC₂H₃O₂, NH₃, KCl
B) HC₂H₃O₂, KCl
C) HCl, NH₃, KCl
D) HCl, HC₂H₃O₂, KCl
E) HCl, KCl
- 4) Which of the following are weak electrolytes? 41 _____
HCl
HC₂H₃O₂
NH₃
KCl
A) HCl, HC₂H₃O₂, NH₃, KCl
B) HCl, KCl
C) HC₂H₃O₂, KCl
D) HC₂H₃O₂, NH₃
E) HCl, HC₂H₃O₂, KCl
- 5) When aqueous solutions of _____ are mixed, a precipitate forms. 51 _____
A) NaI and KBr
B) Li₂CO₃ and CaI
C) K₂SO₄ and CaCl₂
D) KOH and Ba(NO₃)₂
E) NaBr and AgNO₃

Chapter 25 Nuclear Chemistry Workbook Answers

Kenneth W. Whitten, Raymond E. Davis



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