

Chemistry

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(Chapter 1)(Solutions) (Exercises)
(Class 12)

Question 1.26:

If the density of some lake water is 1.25 g mL^{-1} and contains 92 g of Na^+ ions per kg of water, calculate the molality of Na^+ ions in the lake.

Answer 1.26:

Number of moles present in 92 g of Na^+ ions = $\frac{92 \text{ g}}{23 \text{ g mol}^{-1}} = 4 \text{ mol}$

Therefore, molality of Na^+ ions in the lake = $\frac{4 \text{ mol}}{1 \text{ kg}} = 4 \text{ m}$

Question 1.27:

If the solubility product of CuS is 6×10^{-16} , calculate the maximum molarity of CuS in aqueous solution.

Answer 1.27:

Solubility product of CuS , $K_{sp} = 6 \times 10^{-16}$ Let s be the solubility of CuS in mol L^{-1} .



$$K_{sp} = [\text{Cu}^{2+}][\text{S}^{2-}]$$

Now,

$$= s \times s = s^2$$

Then, we have, $K_{sp} = s^2 = 6 \times 10^{-16}$

$$\Rightarrow s = \sqrt{6 \times 10^{-16}} = 2.45 \times 10^{-8} \text{ mol L}^{-1}$$

Hence, the maximum molarity of CuS in an aqueous solution is $2.45 \times 10^{-8} \text{ mol L}^{-1}$.

Question 1.28:

Calculate the mass percentage of aspirin ($\text{C}_9\text{H}_8\text{O}_4$) in acetonitrile (CH_3CN) when 6.5 g of $\text{C}_9\text{H}_8\text{O}_4$ is dissolved in 450 g of CH_3CN .

Answer 1.28:

6.5 g of $\text{C}_9\text{H}_8\text{O}_4$ is dissolved in 450 g of CH_3CN .

Then, total mass of the solution = $(6.5 + 450) \text{ g} = 456.5 \text{ g}$

Therefore, mass percentage of $\text{C}_9\text{H}_8\text{O}_4 = \frac{6.5}{456.5} \times 100 \% = 1.424\%$

Question 1.29:

Nalorphine ($\text{C}_{19}\text{H}_{21}\text{NO}_3$), similar to morphine, is used to combat withdrawal symptoms in narcotic users. Dose of nalorphine generally given is 1.5 mg. Calculate the mass of $1.5 \times 10^{-3} \text{ m}$ aqueous solution required for the above dose.

Answer 1.29:

The molar mass of Nalorphine ($\text{C}_{19}\text{H}_{21}\text{NO}_3$) is given as:

$$19 \times 12 + 21 \times 1 + 1 \times 14 + 3 \times 16 = 311 \text{ g mol}^{-1}$$

In $1.5 \times 10^{-3} \text{ m}$ aqueous solution of nalorphine, 1 kg (1000 g) of water contains $1.5 \times 10^{-3} \text{ mol}$
 $= 1.5 \times 10^{-3} \times 311 \text{ g} = 0.4665 \text{ g}$

Therefore, total mass of the solution = $(1000 + 0.4665) \text{ g} = 1000.4665 \text{ g}$

This implies that the mass of the solution containing 0.4665 g of nalorphine is 1000.4665 g.

Therefore, mass of the solution containing 1.5 mg of nalorphine is: $\frac{1000.4665 \times 1.5 \times 10^{-3}}{0.4665} \text{ g} = 3.22 \text{ g}$

Hence, the mass of aqueous solution required is 3.22 g.

Chapter 1chemistry Answers

Thomas Evangelist



Chapter 1 chemistry Answers:

Solutions Manual for Perspectives on Structure and Mechanism in Organic Chemistry Felix A.

Carroll, 2023-04-11 SOLUTIONS MANUAL FOR PERSPECTIVES ON STRUCTURE AND MECHANISM IN ORGANIC CHEMISTRY Based on the author's first hand classroom experience this solutions manual complements the 3rd edition of Perspectives on Structure and Mechanism in Organic Chemistry The solutions to the 438 textbook problems help students increase their understanding of physical organic chemistry and more than 550 references stimulate their engagement with the chemical literature **General, Organic, and Biochemistry Student's Solutions Manual** Mark D.

Dadmun, 2006-01-06 Provides complete solutions to the odd numbered end of chapter exercises along with additional discussion of problem solving techniques **Wastewater Operator Certification Study Guide** John Giorgi, 2011-01-12

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Transactions of the Pharmaceutical Meetings ,1915 **The Pharmaceutical Journal and Pharmacist** ,1915

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Table of Contents Chapter 1chemistry Answers

1. Understanding the eBook Chapter 1chemistry Answers
 - The Rise of Digital Reading Chapter 1chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Chapter 1chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1chemistry Answers
 - Personalized Recommendations
 - Chapter 1chemistry Answers User Reviews and Ratings

- Chapter 1 chemistry Answers and Bestseller Lists
- 5. Accessing Chapter 1 chemistry Answers Free and Paid eBooks
 - Chapter 1 chemistry Answers Public Domain eBooks
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 - Chapter 1 chemistry Answers Budget-Friendly Options
- 6. Navigating Chapter 1 chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 1 chemistry Answers Compatibility with Devices
 - Chapter 1 chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 1 chemistry Answers
 - Highlighting and Note-Taking Chapter 1 chemistry Answers
 - Interactive Elements Chapter 1 chemistry Answers
- 8. Staying Engaged with Chapter 1 chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 1 chemistry Answers
- 9. Balancing eBooks and Physical Books Chapter 1 chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 1 chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 1 chemistry Answers
 - Setting Reading Goals Chapter 1 chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 1 chemistry Answers
 - Fact-Checking eBook Content of Chapter 1 chemistry Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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