

CHEMISTRY 10TH – Short / Detailed Question Answers

→ CHEMICAL EQUILIBRIUM

CHAPTER# 01

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Q.1: What do you know about chemical reactions?**Ans: Chemical Reactions:**

Many physical and chemical changes are occurring in our surroundings which may be due to chemical reactions. In these reactions, reactants are converted into one or more products.

A chemical reaction is a chemical change which involves reactants and products.

For example : The formation of water from hydrogen gas and oxygen gas, the decomposition reaction of sodium bicarbonate into sodium carbonate, water, and carbon dioxide etc. A chemical reaction contains two quantities; reactant and product which are separated by an arrow.



The reactions may be reversible (condensation, evaporation, freezing, melting) or irreversible (combustion, rusting).

Q.2: Define chemical equilibrium with example**Ans: Chemical Equilibrium:**

The reversible chemical reaction never goes to completion because the product reacts and reproduces reactants again and takes place in forward and reverse directions. This state in which forward reaction rate and reverse reaction rate are equal is known as equilibrium. We know that equilibrium means a balance. Equilibrium exists in many ways in our surroundings.

It is a state where the rate of the forward reaction is equal to the rate of backward reaction OR it is also defined as "A state where the concentration of reactant and product remains constant."

Example: $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons \text{HI}_{(g)}$

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Christian Drosten



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