

MOLE CONCEPT

CLASS Xth

1. Calculate the relative molecular masses (or molecular weights) of the following compounds :

- i) Copper sulphate crystals, $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$.
- ii) Ammonium sulphate, $(\text{NH}_4)_2\text{SO}_4$.
- iii) Cane sugar, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$.

(Given atomic masses Cu = 63.5, S = 32, O = 16, N = 14, C = 12)

2. An organic compound with vapour density = 94 contains. C = 12.67%, H = 2.13%, and Br = 85.11%. Find the molecular formula. [Atomic mass : C = 12, H = 1, Br = 80]

3. (i) The compound A has the following percentage composition by mass. Carbon = 26.7%, oxygen = 71.1%, hydrogen = 2.2%. Determine the Empirical formula of compound A [work to one decimal place] [H = 1, C = 12, O = 16].

ii) If the relative molecular mass of A is 90, what is the molecular formula of A?

(iii) The compound A is weak acid. What is meant by this statement?

4. Determine the empirical formula of the compound whose composition by mass is: 42% nitrogen, 48% oxygen and 9% hydrogen. [H = 1; N = 14; O = 16]

5. Determine the empirical formula of a compound containing 47.9% potassium, 5.5% beryllium and 46.6% fluorine by mass. (Atomic weight of Be = 9; F = 19; K = 39) Work to one decimal place.


6. The Empirical formula of a compound is C_2H_5 . It has a vapour density of 29. Determine the relative molecular formula mass of the compound and hence its molecular formula.

7. The percentage composition of sodium phosphate as determined by analysis, is 42.1% sodium, 18.9% phosphorus and 39% oxygen. Find the empirical formula of the compound (work to two decimal places).

8. A gaseous organic compound contains 3.6 g of carbon and 0.8 g of hydrogen. The vapour density of this compound is 22. (i) Calculate the Empirical formula. (ii) Calculate the molecular formula of the compound.

Chapter 1the Mole Answers

Charles H. Corwin



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Basic Principles of Inorganic Chemistry Brian J Hathaway, Clair Murphy, Brian Murphy, 2007-10-31 General chemistry textbooks are usually lengthy and present chemistry to the student as an unconnected list of facts In inorganic chemistry emphasis should be placed on the connections between valence shell electron configuration and the physical and chemical properties of the element Basic Principles of Inorganic Chemistry Making the Connections is a short concise book that emphasises these connections in particular the chemistry of the Main Group compounds With reference to chemical properties Lewis Structures stoichiometry and spider diagrams students will be able to predict or calculate the chemistry of simple polyatomic compounds from the valence shell configuration and will no longer be required to memorise vast amounts of factual chemistry This book is ideal for students taking chemistry as a subsidiary subject as well as honours degree students [Oswaal JEE \(Main\) Question Bank Chemistry | Chapter-wise & Topic-wise Solved Papers | 2019-2024 | For 2025 Exam](#) Oswaal Editorial Board, 2024-02-28 Description of the Product 100% Updated with 2 latest solved papers of 27th

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Answer Book : Chemistry, a Life Science Approach Stuart J. Baum, C. W. J. Scaife, 1975 **Chemistry Workbook For Dummies** Peter J. Mikulecky, Katherine Brutlag, Michelle Rose Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you ll build your chemistry skills and improve your performance both in and out of the science lab You ll see how to work with numbers atoms and elements make and remake compounds understand changes in terms of energy make sense of organic chemistry and more 100s of Problems Know where to begin and how to solve the most common chemistry problems Step by step answer sets clearly identify where you went wrong or right with a problem

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books photocopiable tests and an answer key *Programmed Problem-solving for First-year Chemistry* Arnold B. Loebel,1983 **Study Guide with Answers to Selected Problems** H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises *Ionic Equilibria in Analytical Chemistry* Henry Freiser,Quintus Fernando,1963 A Study Guide to Chemical Principles Wilbert Hutton,1970 **Living Chemistry** David A. Ucko,1986 **Introductory Chemistry** Charles H. Corwin,2011 To help students learn chemical skills and concepts more effectively Introductory Chemistry Concepts and Critical Thinking Sixth Edition highlights the connection between key concepts and key problem solving skills through critical thinking Math and problem solving are covered early in the text Corwin builds your problem solving ability through innovative learning aids and technology formulated to meet your needs This revision retains all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking Chemical Molecular Science Conrad L. Stanitski,2004-08 How to Prepare for the SAT II Joseph A. Mascetta,1998 An extensive review of chemistry topics a full length diagnostic test precedes four up to date full length model SAT II Chemistry Tests All tests are followed with answers charts students can use to diagnose their individual study needs Chemistry topics reviewed include atomic structure bonding all other topics included in the SAT II test Joseph A Mascetta is a science educational consultant a former principal chemistry teacher at Mount Lebanon High School Pittsburgh PA

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