

Chapter 1

Water in soils

1.1 Steady-state fluid flow

Exercise 1

With reference to the permeameter in Figure 1.1, define the pore pressure in Y if 35% of the difference bw. the hydraulic head values dissipates in soil A (Note: lengths in cm). If the width of the permeameter is 100cm and the hydraulic conductivity K_B of soil B is $1 \times 10^{-2} \text{ cm/s}$, define also the discharge Q in m^3/s crossing the permeameter and the hydraulic conductivity K_A of soil A.

Solution

A permeameter, in the original setup of Darcy (see Figure 1.2), is a laboratory tool for the measurement of the hydraulic conductivity (o hydrogeological permeability) K of a given soil. It is a cylinder where a soil specimen is inserted; at the top and at the bottom there are two porous stones acting as filters (the water particles pass through, whereas the solid particles are kept in). The stones are connected to two reservoirs, whose levels are h_1 and h_2 (with respect to the base in the figure), maintained fixed by feeding the soil with water to produce a downward steady-state flow.

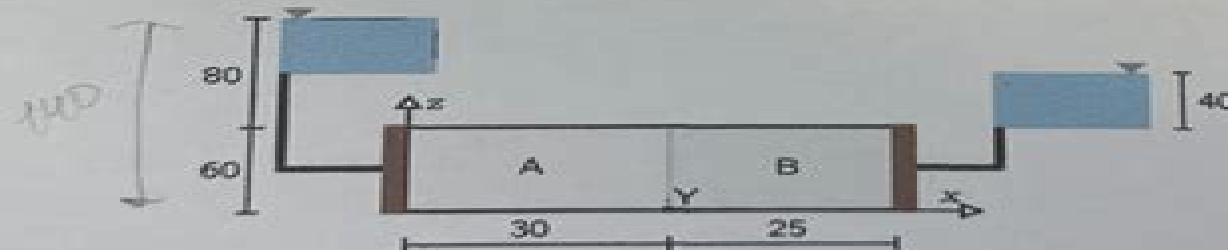


Figure 1.1. Scheme exercise #1

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Thermodynamics of Solutions Eli Ruckenstein, Ivan L. Shulgin, 2009-06-17 This book consists of a number of papers regarding the thermodynamics and structure of multicomponent systems that we have published during the last decade. Even though they involve different topics and different systems, they have something in common which can be considered as the signature of the present book. First, these papers are concerned with difficult or very nonideal systems, i.e., systems with very strong interactions, e.g., hydrogen bonding between components or systems with large differences in the partial molar volumes of the components, e.g., the aqueous solutions of proteins or systems that are far from normal conditions, e.g., critical or near-critical mixtures. Second, the conventional thermodynamic methods are not sufficient for the accurate treatment of these mixtures. Last but not least, these systems are of interest for the pharmaceutical, biomedical, and related industries. In order to meet the thermodynamic challenges involved in these complex mixtures, we employed a variety of traditional methods but also new methods such as the fluctuation theory of Kirkwood and Buff and *ab initio* quantum mechanical techniques. The Kirkwood-Buff (KB) theory is a rigorous formalism which is free of any of the approximations usually used in the thermodynamic treatment of multicomponent systems. This theory appears to be very fruitful when applied to the above-mentioned difficult systems.

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